

**1. The whoosh jug**

## a. Observations:

When a match is dropped into the jug, a big whoosh of flame.

b. Reaction for combustion of ethyl alcohol ( $\text{C}_2\text{H}_5\text{OH}$ ):

## c. What is it that is whooshing out of the jug?

When the burn occurs, more gases are formed than are started with. AND the gases get heated up which make them expand. Both of these reactions cause gases to push up and out of the top of the jug.

## d. Why was the jug rotated and twirled for a minute or two before the reaction?

The twirling was necessary to help the alcohol to evaporate and turn into a vapor.

## e. Extra unburned liquid alcohol remained in the jug after the initial combustion. This alcohol could be poured out onto the lab bench and ignited. Why did it not all burn inside the jug?

There was no oxygen left in the jug after the first burn thus a second combustion could not occur in the jug, just at the mouth of the jug.

## f. When the same alcohol is ignited on the lab bench, why doesn't it give off a big whoosh?

The whoosh will only occur when the burn takes place in an enclosed container, and the hot expanding gases push out through the small opening.

## g. Why does the jug collapse when the opening was covered right after the whoosh?

The hot gases all whooshed out. The gases left are hot, but as soon as the flame was extinguished, the gases began to cool and shrink (or cause less pressure). Since the opening is covered, the air can't push back in through the opening, so the air tries to push in through the plastic, collapsing the jug.

**2. Burning dollars**

## a. Observations:

Dipped in one liquid the paper (dollar) does not burn. Dipped in the other solution, the paper burns.

b. Reaction for combustion of isopropyl alcohol ( $\text{C}_3\text{H}_8\text{OH}$ ):

## c. Why did the paper burn, yet the dollar bill remained intact?

The bill was dipped in a 50% alcohol/50% water solution. The presence of the water absorbed much of the energy from the burning alcohol so that the bill never got hot enough to ignite.

## d. Why does the fire department use water to put out fires?

Water is used to put out fires because it cools the unburned fuel (wood that hasn't burned yet) so much that there isn't enough heat to allow the combustion to continue. For fire to occur, all three items in the fire triangle must be present.

**3. Methane bubbles vs Propane bubbles**

## a. Observations:

Methane bubbles rise to the ceiling, and propane bubbles sink to the floor. Both bubbles ignited when lit with the candle.

b. Reaction for combustion of propane ( $\text{C}_3\text{H}_8$ ):c. Methane is  $\text{CH}_4$ , propane is  $\text{C}_3\text{H}_8$ , oxygen is  $\text{O}_2$ , and nitrogen is  $\text{N}_2$ . What do the molar masses of these chemicals imply about the density of these gases? Why does the methane balloon rise up to the ceiling, yet the propane bubble sinks to the floor?

Methane,  $\text{CH}_4$  weighs 16 g/mole, propane,  $\text{C}_3\text{H}_8$  weighs 44 g/mole. Air weighs 29 g/mole (80%  $\text{N}_2 = 28$ , 20%  $\text{O}_2 = 32$ ). This methane is less dense than air so it floats, and propane is more dense than air so it sinks.

# Blue Devil Chem Ap Gases

**SJ Ball**



**Blue Devil Chem Ap Gases:**

## **Blue Devil Chem Ap Gases Book Review: Unveiling the Power of Words**

In a global driven by information and connectivity, the power of words has are more evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Blue Devil Chem Ap Gases**, a literary masterpiece that delves deep to the significance of words and their impact on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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