

Boyles law questions

SCH3U Boyle's Law Worksheet

1. State the pressure-volume law both in words and in the form of an equation.
2. To compress nitrogen at 1 atm from 750 mL to 500 mL, what must the new pressure be if the temperature is kept constant? **1.5 atm**
3. If oxygen at 128 kPa is allowed to expand at constant temperature until its pressure is 101.3 kPa, how much larger will the volume become? **1.26**
4. A sample of nitrogen at 101.3 kPa with a volume of 100 mL is carefully compressed at constant temperature in successive changes in pressure, equalling 5 kPa at a time, until the final pressure is 133.3 kPa. Calculate each new volume and prepare a plot of P versus V, showing P on the horizontal axis.
5. A sample of nitrogen at 20°C was compressed from 300 mL to 0.360 mL and its new pressure was found to be 400.0 Pa. What was the original pressure in kPa? **4.8×10^{-4}**
6. The pressure on 6.0 L of a gas is 200 kPa. What will be the volume if the pressure is doubled, keeping the temperature constant? **3L**
7. What would be the new volume if the pressure on 600 mL is increased from 90 kPa to 150 kPa?
360 mL
8. A student collects 25 mL of gas at 96 kPa. What volume would this gas occupy at 101.325 kPa. There is no change in temperature or mass. **23.68 mL**
9. A gas measuring 525 mL is collected at 104.66 kPa. What volume does this gas occupy at 99.33 kPa? **553.17 mL**
10. A mass of gas occupies 1 L at 1 atm. At what pressure does this gas occupy
a) 2 litres, **0.5 atm**
b) 0.5 litres? **2 atm**
11. From the data in the following table calculate the missing quantity (assuming constant temperature).
a) $V_1 = 22.4 \text{ L}$; $P_1 = 1 \text{ atm}$; $P_2 = ? \text{ atm}$; $V_2 = 2.8 \text{ L}$ **8 atm**
b) $V_1 = 60 \text{ mL}$; $P_1 = ? \text{ kPa}$; $P_2 = 101.3 \text{ kPa}$; $V_2 = 16 \text{ mL}$ **27 kPa**
c) $V_1 = ? \text{ m}^3$; $P_1 = 40 \text{ Pa}$; $P_2 = 100 \text{ kPa}$; $V_2 = 1.0 \text{ L}$ **$2.5 \times 10^9 \text{ m}^3$**
d) $V_1 = 2.50 \text{ L}$; $P_1 = 7.5 \text{ atm}$; $P_2 = ? \text{ atm}$; $V_2 = 100 \text{ mL}$ **187.5 atm**

Boyles Law Cavalcade Publishing Answers

A Gutmann



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