

Chemistry 338 2nd Semester Final Exam Review Sheet

Unit 1 - Matter & Measurement – Chapters 1 & 2

1. A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

$$\% \text{ error} = \frac{\text{experimental} - \text{accepted}}{\text{accepted}} \times 100$$

$$\% \text{ error} = \frac{4.2 - 5.0}{5.0} \times 100 = -16\% \text{ error}$$

Unit 2 – Periodic Table – Chapter 5

2.

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten ion	¹⁸⁴ ₇₄ W ⁺²	74	74	110	72	184	+2

3. Using the following atomic radius values, predict the atomic radius of potassium (K):
 Li – 145 pm, Na – 180 pm, K - ?, Rb – 235 pm
K – 220 pm

Unit 3 – Reactions/Compounds – Chapters 7 & 8

4. Classify the compound as ionic or covalent. Name the compound
- Fe₂O₃ - iron (III) oxide (ionic - roman numerals for transition metals, change ending to -ide)
 - Cl₂O₈ - dichlorine octaoxide (covalent - use prefixes, change ending to -ide)
5. Classify the compound as ionic or covalent. Write the formula
- Trisulfur Dinitride – S₃N₂ (covalent)
 - Zinc Hydroxide – Zn(OH)₂ (ionic – look up ion charges and criss-cross) Zn⁺² OH⁻¹
 - Chromium (III) nitrate – Cr(NO₃)₃ (ionic – criss-cross) Cr⁺³ NO₃⁻²
6. Write the equation, balance it, and classify the type of reaction:
 Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate
- $$1 \text{ FeSO}_4 + 2 \text{ NaOH} \rightarrow 1 \text{ Fe(OH)}_2 + 1 \text{ Na}_2\text{SO}_4$$
- double displacement (two ionic compounds reacting)
7. Balance the equation and classify the type of reaction:
- $$1 \text{ Cu} + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Cu(NO}_3)_2 + 2 \text{ Ag}$$
- Single displacement (element and a compound reacting)

Unit 4 – Moles – Chapter 3

8. Calculate the molar mass of Fe₂(SO₄)₃.

$$\begin{aligned} \text{Fe (2)(55.9 g/mol)} &= 111.8 \text{ g/mol} \\ \text{S (3)(32.1 g/mol)} &= 96.3 \text{ g/mol} \\ + \text{O (12)(16.0 g/mol)} &= +192 \text{ g/mol} \\ &= 400.1 \text{ g/mol} \end{aligned}$$

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Michael Brown



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Table of Contents Nd Semester Final Review Guide Chemistry

1. Understanding the eBook Nd Semester Final Review Guide Chemistry
 - The Rise of Digital Reading Nd Semester Final Review Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Nd Semester Final Review Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nd Semester Final Review Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nd Semester Final Review Guide Chemistry
 - Personalized Recommendations
 - Nd Semester Final Review Guide Chemistry User Reviews and Ratings
 - Nd Semester Final Review Guide Chemistry and Bestseller Lists
5. Accessing Nd Semester Final Review Guide Chemistry Free and Paid eBooks
 - Nd Semester Final Review Guide Chemistry Public Domain eBooks
 - Nd Semester Final Review Guide Chemistry eBook Subscription Services
 - Nd Semester Final Review Guide Chemistry Budget-Friendly Options

6. Navigating Nd Semester Final Review Guide Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Nd Semester Final Review Guide Chemistry Compatibility with Devices
 - Nd Semester Final Review Guide Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nd Semester Final Review Guide Chemistry
 - Highlighting and Note-Taking Nd Semester Final Review Guide Chemistry
 - Interactive Elements Nd Semester Final Review Guide Chemistry
8. Staying Engaged with Nd Semester Final Review Guide Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nd Semester Final Review Guide Chemistry
9. Balancing eBooks and Physical Books Nd Semester Final Review Guide Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nd Semester Final Review Guide Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nd Semester Final Review Guide Chemistry
 - Setting Reading Goals Nd Semester Final Review Guide Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nd Semester Final Review Guide Chemistry
 - Fact-Checking eBook Content of Nd Semester Final Review Guide Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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