Name: Date:

Limiting Reagent Worksheet

Answers

Given the following reaction:

$$C_1H_A + \underline{5} O_1 \rightarrow \underline{3} CO_1 + \underline{4} H_1O_2$$

Balance the reaction.

Suppose 14.8 g of C₃H₈ reacts with 3.44 g of O₂. Determine the limiting reagent.

$$14.8 \text{ g C}_{3}H_{8} \times \frac{1 \text{ mol C}_{3}H_{8}}{44.09 \text{ g C}_{3}H_{8}} \times \frac{4 \text{ mol H}_{2}O}{1 \text{ mol C}_{3}H_{8}} \times \frac{18.02 \text{ g H}_{2}O}{1 \text{ mol H}_{2}O} = 24.2 \text{ g H}_{2}O$$

$$3.44 \text{ g O}_{2} \times \frac{1 \text{ mol O}_{2}}{32 \text{ g O}_{2}} \times \frac{4 \text{ mol H}_{2}O}{5 \text{ mol O}_{2}} \times \frac{18.02 \text{ g H}_{2}O}{1 \text{ mol H}_{2}O} = 1.55 \text{ g H}_{2}O$$

Since O, produces less quantity of H,O than C,H, does, it is the limiting reagent.

(b) Determine the amount of CO, produced.

2 Given the following reaction:

Balance the reaction.

What is the limiting reagent if 2.2 g of Mg reacts with 4.5 L of O, at STP?

2.2 g Mg *
$$\frac{1 \text{ mol Mg}}{24.3 \text{ g Mg}}$$
 * $\frac{2 \text{ mol MgO}}{2 \text{ mol Mg}}$ = 0.0905 mol MgO
4.5 L O₂ * $\frac{1 \text{ mol O}_2}{22.4 \text{ L O}_2}$ * $\frac{2 \text{ mol MgO}}{1 \text{ mol O}_2}$ * = 0.401 mol MgO

Since Mg produces less MgO than O, Mg is the limiting reagent.

Determine the amount of MgO produced in grams.

Limiting Reagent Worksheet 202 Answer Key

Hussin A.Rothana

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