



## Limiting Reactants and Theoretical Yield, continued

### The Limiting Reactant Forms the Least Product

- Reactants of a reaction are seldom present in ratios equal to the mole ratio in the balanced equation, so one of the reactants is used up first.
  - For example, If you combine 0.23 mol Zn and 0.60 mol HCl, would they react completely?



- The coefficients indicate that 0.23 mol Zn forms 0.23 mol  $\text{H}_2$ , and 0.60 mol HCl forms 0.30 mol  $\text{H}_2$ .



# Limiting Reactants And Percentage Yield Concept Review

**Barron's Educational Series, Neil D.  
Jespersen, Pamela Kerrigan**



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