

Limiting Reactants & Excess

- The **limiting reactant** is the reactant in the smallest amount according to the mole ratio.
- Whenever the quantities of both reactants are given, the limiting reactant must be found.
- To find the limiting reactant, convert quantities of both reactants into moles and compare with the mole ratio.
 - An **excess** reactant will be higher than the amount required by ratio.

- **Example**



$$m(\text{Mg}) = 1.34 \text{ g}; c(\text{HCl}) = 1.00 \text{ mol L}^{-1}; V(\text{HCl}) = 0.0600 \text{ L}$$

$$n(\text{Mg}) = \frac{1.34}{24.3} = 0.0551 \text{ mol}; n(\text{HCl}) = 1.00 \times 0.0600 = 0.0600 \text{ mol}$$

The mole ratio should be 1 : 2, i.e. 0.0551 : 0.110, but there is insufficient HCl, making it the limiting reactant used in further calculations.

Limiting And Excess Reactants Chembridge Pogil

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