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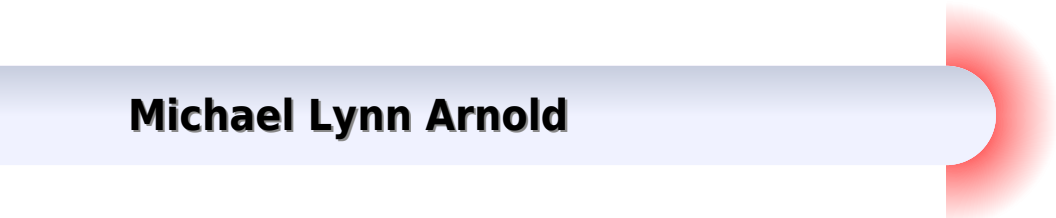
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MARKS: 40

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temperature rose from 23.42 °C to 27.64 °C and the heat capacity of the calorimeter is 4.90 kJ/°C then determine the heat of combustion of sucrose $C_{12}H_{22}O_{11}$ in kJ per mole of $C_{12}H_{22}O_{11}$

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web in a calorimeter the white solid $CaSO_4$ forms and the temperature of the mixture increases to 28.1 °C assuming that the calorimeter absorbs only a negligible quantity of heat and that the specific heat capacity of the solution is 4.18 J/°C·g and that the density of the final solution is 1.0 g/mL calculate the enthalpy change per mole

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