<u>Limiting Reagent Worksheet #1</u>

1. Given the following reaction: (Balance the equation first!)

$$C_3H_8 + O_2 -----> CO_2 + H_2O$$

- a) If you start with 14.8 g of C3H8 and 3.44 g of O2, determine the limiting reagent
- b) determine the number of moles of carbon dioxide produced
- c) determine the number of grams of H2O produced
- d) determine the number of grams of excess reagent left
- 2. Given the following equation:

$$Al_2(SO_3)_3 + 6 NaOH ----> 3 Na_2SO_3 + 2 Al(OH)_3$$

- a) If 10.0 g of Al₂(SO₃)₃ is reacted with 10.0 g of NaOH, determine the limiting reagent
- b) Determine the number of moles of Al(OH)₃ produced
- c) Determine the number of grams of Na₂SO₃ produced
- d) Determine the number of grams of excess reagent left over in the reaction
- 3. Given the following equation:

$$Al_2O_3 + Fe ----> Fe_3O_4 + Al$$

- a) If 25.4 g of Al₂O₃ is reacted with 10.2 g of Fe, determine the limiting reagent
- b) Determine the number of moles of Al produced
- c) Determine the number of grams of Fe₃O₄ produced
- d) Determine the number of grams of excess reagent left over in the reaction

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