

Substances held together by covalent bonds are usually divided into two groups based on whether individual (distinct) molecules exist or not. In a *molecular solid*, individual molecules in the solid state are attracted to each other by relatively weak intermolecular forces between the molecules. *Covalent-network solids*, on the other hand, consist of atoms forming covalent bonds with each other in all directions. The result is an almost infinite network of strong covalent bonds—there are no individual molecules.

Covalent bonds may be classified as polar or nonpolar. The element chlorine, for example, exists as a diatomic molecule, Cl_2 . The two chlorine atoms are held together by a single covalent bond, with the two electrons in the bond equally shared between the two identical chlorine atoms. This type of bond is called a *nonpolar* covalent bond. The compound hydrogen chloride (HCl) consists of a hydrogen atom and a chlorine atom that also share a pair of electrons between them. Because the two atoms are different, however, the electrons in the bond are not equally shared between the atoms. Chlorine has a greater *electronegativity* than hydrogen—it attracts the bonding electrons more strongly than hydrogen. The covalent bond between hydrogen and chlorine is an example of a *polar* bond. The distribution of bonding electrons in a nonpolar versus polar bond is shown in Figure 2. Notice that the chlorine atom in HCl has a partial negative charge (δ^-) while the hydrogen atom has a partial positive charge (δ^+).

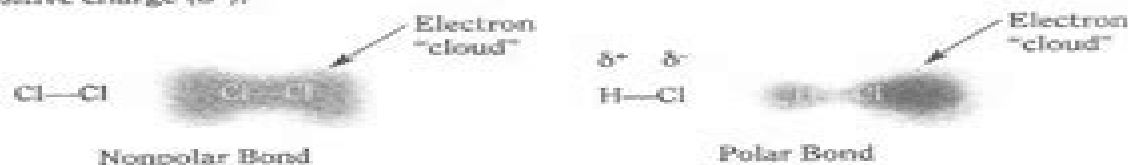


Figure 2. Nonpolar versus Polar Covalent Bonds.

The special properties of metals compared to nonmetals reflect their unique structure and bonding. Metals typically have a small number of valence electrons available for bonding. The valence electrons appear to be free to move among all of the metal atoms, which must exist therefore as positively charged cations. *Metallic bonding* describes the attractive forces that exist between closely packed metal cations and free-floating valence electrons in an extended three-dimensional structure.

Experiment Overview

The purpose of this experiment is to study the physical properties of common solids and to investigate the relationship between the type of bonding in a substance and its properties. The following physical properties will be studied:

- **Volatility and Odor:** Volatile substances evaporate easily and may have an odor.
- **Melting Point:** The temperature at which a solid turns into a liquid.
- **Solubility:** Ability of one substance to dissolve in another. Water is a highly polar solvent. Hexane is nonpolar.
- **Conductivity:** Ability to conduct electricity.
- **Hardness:** Resistance of a substance to being scratched.
- **Brittleness:** Tendency of a solid to break or crumble when a stress is applied.

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R Bogdan



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Table of Contents D3jc3ahdjad7x7 Cloudfront Mitosis

1. Understanding the eBook D3jc3ahdjad7x7 Cloudfront Mitosis
 - The Rise of Digital Reading D3jc3ahdjad7x7 Cloudfront Mitosis
 - Advantages of eBooks Over Traditional Books
2. Identifying D3jc3ahdjad7x7 Cloudfront Mitosis
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an D3jc3ahdjad7x7 Cloudfront Mitosis
 - User-Friendly Interface
4. Exploring eBook Recommendations from D3jc3ahdjad7x7 Cloudfront Mitosis

- Personalized Recommendations
- D3jc3ahdjad7x7 Cloudfront Mitosis User Reviews and Ratings
- D3jc3ahdjad7x7 Cloudfront Mitosis and Bestseller Lists
- 5. Accessing D3jc3ahdjad7x7 Cloudfront Mitosis Free and Paid eBooks
 - D3jc3ahdjad7x7 Cloudfront Mitosis Public Domain eBooks
 - D3jc3ahdjad7x7 Cloudfront Mitosis eBook Subscription Services
 - D3jc3ahdjad7x7 Cloudfront Mitosis Budget-Friendly Options
- 6. Navigating D3jc3ahdjad7x7 Cloudfront Mitosis eBook Formats
 - ePub, PDF, MOBI, and More
 - D3jc3ahdjad7x7 Cloudfront Mitosis Compatibility with Devices
 - D3jc3ahdjad7x7 Cloudfront Mitosis Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of D3jc3ahdjad7x7 Cloudfront Mitosis
 - Highlighting and Note-Taking D3jc3ahdjad7x7 Cloudfront Mitosis
 - Interactive Elements D3jc3ahdjad7x7 Cloudfront Mitosis
- 8. Staying Engaged with D3jc3ahdjad7x7 Cloudfront Mitosis
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers D3jc3ahdjad7x7 Cloudfront Mitosis
- 9. Balancing eBooks and Physical Books D3jc3ahdjad7x7 Cloudfront Mitosis
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection D3jc3ahdjad7x7 Cloudfront Mitosis
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine D3jc3ahdjad7x7 Cloudfront Mitosis
 - Setting Reading Goals D3jc3ahdjad7x7 Cloudfront Mitosis
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of D3jc3ahdjad7x7 Cloudfront Mitosis

- Fact-Checking eBook Content of D3jc3ahdjad7x7 Cloudfront Mitosis
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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