

Name: _____

Naming Practice

When dealing with compounds, the first thing that should be done is determine if they are ionic or covalent. And if I am having you name them and it has a metal, then it is ionic. If you see nonmetals then it is covalent.

Dealing with Ionic Compounds

Since ionic compounds have ions, these are the ones that you have to worry about charges with. Most of the time you can use the periodic table to determine the charge. If the element is from the first column of the periodic table, it has a +1 charge. Second column = +2 charge. Group 13 (Boron Family) = +3 charge. Group 15 = -3 charge. Group 16 = -2 charge. Group 17 = -1 charge. If it is a transition element, inner transition, or Group 14, we will assume they have more than one charge, and we will say what charge it is by using a roman numeral.

Ionic Compounds also deal more with the polyatomic ions. If you see -ate or -ite in the name, you are dealing with a polyatomic. If you are dealing with an element, you will use the ending -ide. The tricky polyatomic ions are ammonium (NH_4^+), cyanide (CN^-), hydroxide (OH^-), and peroxide (O_2^{2-}) (Their formulas go in the blanks). The final thing you have to worry about is switchy-switchy, which is switching the charge of one ion to be the subscript of the other part of the compound, and reduce the numbers if possible. Remember, you only take the number and not the sign. Let's practice:

Compound Name	Ions that make up the compound	Resulting Formula
calcium chloride	$\text{Ca}^{+2} \text{Cl}^{-1}$	CaCl_2
cobalt (V) nitride	$\text{Co}^{+5} \text{N}^{-3}$	Co_3N_5
strontium hydroxide	$\text{Sr}^{+2} (\text{OH})^{-1}$	$\text{Sr}(\text{OH})_2$
lead (III) sulfate	$\text{Pb}^{+3} (\text{SO}_4)^{-2}$	$\text{Pb}_2(\text{SO}_4)_3$
iron (III) phosphate	$\text{Fe}^{+3} (\text{PO}_4)^{-3}$	FePO_4 (reduces)
magnesium fluoride	$\text{Mg}^{+2} \text{F}^{-1}$	MgF_2
chromium (III) sulfite	$\text{Cr}^{+3} (\text{SO}_3)^{-2}$	$\text{Cr}_2(\text{SO}_3)_3$
barium selenide	$\text{Ba}^{+2} \text{Se}^{-2}$	BaSe
osmium (IV) oxide	$\text{Os}^{+4} \text{O}^{-2}$	OsO_2

reverse switchy-switchy doesn't work on these two.
start w/ anion 1st (from P.T.)

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Peck Ritter



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