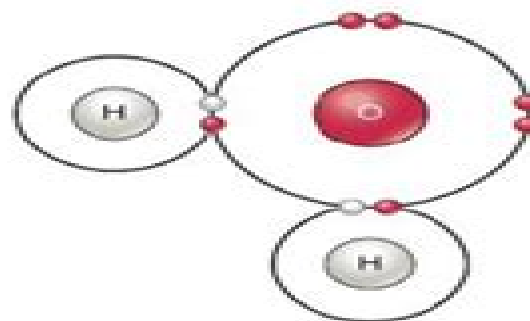
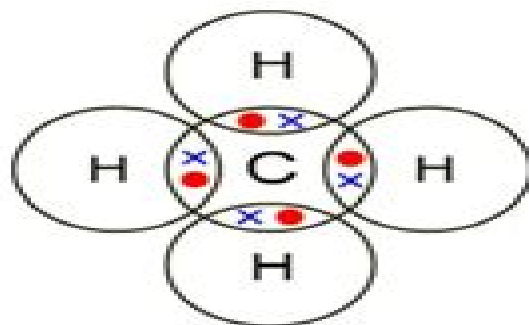
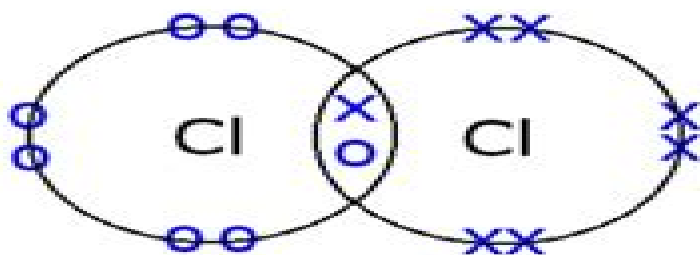


## COVALENT BONDING

- Chemical bond formed by sharing of electrons (from the valence shell) between two atoms, so that each atom in the resulting molecule gets the stable electronic configuration of the nearest noble gas.
- A covalent bond is formed only when both reacting atoms need electrons to achieve stable electronic configuration. ie, Non-metal atoms form covalent bonds
- When two non-metal atoms come close enough, their shells overlap and they share electrons. The strong force of attraction developed between the atoms when they share electrons is called covalent bond.
- Three types of covalent bond – Single , Double , Triple – based on the number of electrons shared by each non-metal atom pair.



# Covalent Bond Notes Ch 8

**R Barnett**

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