

Chemistry Chapter 6

- Real Gases
 - Real gases often do not behave like ideal gases at high pressure or low temperature.
 - Ideal gas laws assume
 - 1. no attractions between gas molecules.
 - 2. gas molecules do not take up space.
 - Based on the kinetic molecular theory
 - At low temperatures and high pressures these assumptions are not valid.
- The Effect of the Finite Volume of Gas Particles
 - At low pressures, the molar volume of argon is nearly identical to that of an ideal gas.
 - But as the pressure increases, the molar volume of argon becomes greater than that of an ideal gas.
 - At the higher pressures, there are more argon atoms. The volume occupied by these atoms is not negligible.
- Modification of the Ideal Gas Equation
 - In 1873, Johannes van der Waals (1837–1923) modified the ideal gas equation to fit the behavior of real gases at high pressure.
 - The molecular volume makes the real volume larger than the ideal gas law would predict.
 - Van der Waals modified the ideal gas equation to account for the molecular volume.
 - b is called a van der Waals constant and is different for every gas because their molecules are different sizes.
- The Effect of Intermolecular Attractions
 - At high temperature, the pressures of the gases are nearly identical to that of an ideal gas.
 - But at lower temperatures, the pressures of the gases are less than that of an ideal gas.
 - i. At the lower temperatures, the gas atoms spend more time interacting with each other and less time colliding with the walls, making the actual pressure less than that predicted by the ideal gas law.
- The Effect of Intermolecular Attractions
 - Van der Waals modified the ideal gas equation to account for the intermolecular attractions.
 - i. a is another van der Waals constant and is different for every gas because their molecules have different strengths of attraction.
- Properties of Gases
 - Expand to completely fill their container
 - Take the shape of their container
 - Low density, much less than solid or liquid state

Core Teaching Resources Chemistry Chapter 6

DP Hallahan



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