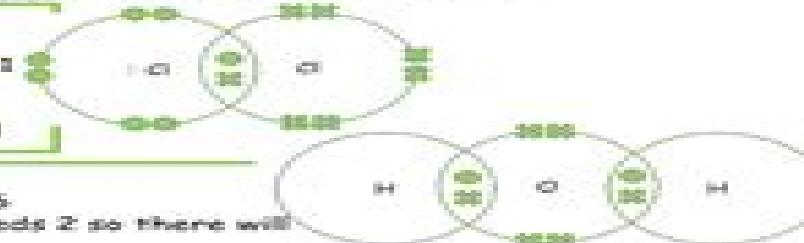


# AQA Science - Covalent bonding

- Identify the group numbers of the elements involved
- The number of electrons needed to become stable is the number of 'pairs' in the shared area (each pair should be made of one dot and one cross)
- Fill in the shared area with the pairs
- The electrons remaining from the original atom are placed on the outer shell.

Worked example 1: Chlorine -  $\text{Cl}_2$

- Chlorine is in group 7
- Chlorine needs 1 electron to become stable so there will be 1 pair of electrons in the shared area
- Fill in the shared area with the 1 pair
- There are 6 electrons left so place these on the outer shell



Worked example 2: water  $\text{H}_2\text{O}$

- Hydrogen has 1 electron in the 1<sup>st</sup> shell. Oxygen is in group 6
- Hydrogen needs 1 electron to become stable and oxygen needs 2 so there will be 2 hydrogen atoms
- Fill in the shared areas with 1 pair in each
- The only hydrogen electron is used so there are no others to place. There are 4 electrons left on the oxygen atom and so these are placed on the outer shell

## Your turn

Hydrochloric acid -  $\text{HCl}$

Ammonia -  $\text{NH}_3$

Methane -  $\text{CH}_4$

Carbon Dioxide -  $\text{CO}_2$

Oxygen -  $\text{O}_2$

Nitrogen -  $\text{N}_2$



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**G Thomas**

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