## QUESTION 10 (Start on a new page.)

The balanced chemical equation for the EXTRACTION of gold from its ore is as follows:

$$4Au(s) + 8NaCN(aq) + 2H2O(l) + O2(g) \rightarrow 4NaAu(CN)2(aq) + 4NaOH(aq)$$

- 10.1 State ONE disadvantage of using cyanide (CN') in the extraction of gold. (1)
- 10.2 Will the final solution of the extraction process be ACIDIC or BASIC (ALKALINE)? Give a reason for the answer.
  (2)
- 10.3 Determine the oxidation number of gold in NaAu(CN)<sub>2</sub>. (1)
- 10.4 Write down the FORMULA of the reducing agent in the reaction above. (1)

Zinc powder is now used to PRECIPITATE the gold.

The balanced equation for the reaction is:

$$Zn(s) + 2NaAu(CN)_2(aq) \rightarrow 2Au(s) + Zn(CN)_2(aq) + 2NaCN(aq)$$

- 10.5 Does zinc undergo OXIDATION or REDUCTION during the precipitation reaction? (1)
- 10.6 Write down a half-reaction to support the answer to QUESTION 10.5. (2)
- 10.7 Calculate the percentage of gold in NaAu(CN)<sub>2</sub>. (2)
  [10]

TOTAL: 150

## **Dbe November 2014 Physical Sciences P2**

Paul G. Hewitt, John A. Suchocki, Leslie A. Hewitt

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- 1. Understanding the eBook Dbe November 2014 Physical Sciences P2
  - The Rise of Digital Reading Dbe November 2014 Physical Sciences P2
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Dbe November 2014 Physical Sciences P2
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Dbe November 2014 Physical Sciences P2
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Dbe November 2014 Physical Sciences P2
  - Personalized Recommendations
  - Dbe November 2014 Physical Sciences P2 User Reviews and Ratings
  - Dbe November 2014 Physical Sciences P2 and Bestseller Lists
- 5. Accessing Dbe November 2014 Physical Sciences P2 Free and Paid eBooks
  - Dbe November 2014 Physical Sciences P2 Public Domain eBooks
  - Dbe November 2014 Physical Sciences P2 eBook Subscription Services
  - Dbe November 2014 Physical Sciences P2 Budget-Friendly Options
- 6. Navigating Dbe November 2014 Physical Sciences P2 eBook Formats
  - o ePub, PDF, MOBI, and More
  - Dbe November 2014 Physical Sciences P2 Compatibility with Devices
  - Dbe November 2014 Physical Sciences P2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Dbe November 2014 Physical Sciences P2
  - Highlighting and Note-Taking Dbe November 2014 Physical Sciences P2
  - Interactive Elements Dbe November 2014 Physical Sciences P2
- 8. Staying Engaged with Dbe November 2014 Physical Sciences P2

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Dbe November 2014 Physical Sciences P2
- 9. Balancing eBooks and Physical Books Dbe November 2014 Physical Sciences P2
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Dbe November 2014 Physical Sciences P2
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Dbe November 2014 Physical Sciences P2
  - Setting Reading Goals Dbe November 2014 Physical Sciences P2
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Dbe November 2014 Physical Sciences P2
  - Fact-Checking eBook Content of Dbe November 2014 Physical Sciences P2
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - $\circ \ \ Integration \ of \ Multimedia \ Elements$
  - Interactive and Gamified eBooks

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