

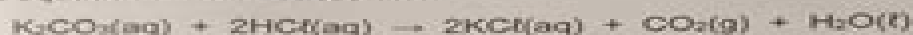
QUESTION 7 (Start on a new page.)

Hydrated potassium carbonate, $K_2CO_3 \cdot xH_2O$, is a WEAK BASE. A solution is prepared by dissolving some of this solid in water.

- 7.1 Define the term weak base. (2)
- 7.2 Write down the formula of the conjugate acid of the carbonate ion, $CO_3^{2-}(aq)$. (1)

A hydrochloric acid solution, $HCl(aq)$, of concentration $0,1 \text{ mol} \cdot \text{dm}^{-3}$ is titrated with the prepared potassium carbonate solution, $K_2CO_3(aq)$, of unknown concentration.

The balanced equation for the reaction is:



The results of the titration are given below.

	VOLUME OF $HCl(aq)$ USED (cm^3)	$K_2CO_3(aq)$ IN BURETTE		VOLUME OF $K_2CO_3(aq)$ USED (cm^3)
		INITIAL BURETTE READING (cm^3)	FINAL BURETTE READING (cm^3)	
Run 1	25	6,5	p	20,05
Run 2	25	q	48,3	20,15

- 7.3 Determine the value of:
- 7.3.1 p (1)
- 7.3.2 q (1)
- 7.4 METHYL ORANGE is used as the indicator. Explain why methyl orange is the most suitable indicator for this titration by referring to the pH at the equivalence point. (2)
- 7.5 Calculate the concentration of the K_2CO_3 solution. (5)
- The above K_2CO_3 solution used in the titration, was prepared by completely dissolving 6,525 g of the hydrated potassium carbonate, $K_2CO_3 \cdot xH_2O$, in 600 cm^3 water.
- 7.6 Calculate the value of x in the formula $K_2CO_3 \cdot xH_2O$. (5)

[17]



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