

COVALENT BONDING

Name _____

Covalent bonding occurs when two or more nonmetals share electrons, attempting to attain a stable octet of electrons at least part of the time. For example:

*H wants 2 electrons
all other atoms 8*



electrons in the middle share

Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

wanted available share

<p>1. H + H (H₂)</p> <p>W - 4 A - 2 S = 2</p>	<p><i>wanted 2 + 2 available</i></p> <p>$\text{H} \cdot \cdot \text{H}$</p> <p>$\text{H} \cdot \quad \cdot \text{H}$</p>	<p>$\text{H}-\text{H}$</p> <p>single bond = share 2 e⁻</p>
<p>2. F + F (F₂)</p> <p>W - 8 + 8 = 16 A - 7 + 7 = 14 S = 2</p>	<p>$:\ddot{\text{F}}:\ddot{\text{F}}:$</p> <p>$\text{F} \cdot \cdot \cdot \text{F}$</p>	<p>$:\ddot{\text{F}}-\ddot{\text{F}}:$</p> <p>single bond 2e⁻</p>
<p>3. O + O (O₂)</p> <p>W - 8 + 8 = 16 A - 6 + 6 = 12 S = 4</p>	<p>$:\ddot{\text{O}}::\ddot{\text{O}}:$</p> <p>$\text{O} \cdot \cdot \cdot \cdot \text{O}$</p>	<p>$:\ddot{\text{O}}=\ddot{\text{O}}:$</p> <p>4 double bond</p>
<p>4. N + N (N₂)</p> <p>W - 8 + 8 = 16 A - 5 + 5 = 10 S = 6</p>	<p>$:\text{N}:::\text{N}:$</p> <p>$\text{N} \cdot \cdot \cdot \cdot \cdot \text{N}$</p>	<p>$:\text{N}\equiv\text{N}:$</p>
<p>5. C + O (CO₂)</p> <p>W - 8 + 8 + 8 = 24 A - 4 + 6 + 6 = 16 S = 8</p>	<p>$:\ddot{\text{O}}:::\text{C}:::\ddot{\text{O}}:$</p> <p>$\text{O} \cdot \cdot \cdot \cdot \cdot \text{C} \cdot \cdot \cdot \cdot \cdot \text{O}$</p>	<p>$\ddot{\text{O}}=\text{C}=\ddot{\text{O}}$</p>
<p>6. H + O (H₂O)</p> <p>W - 2 + 2 + 8 = 12 A - 1 + 1 + 6 = 8 S = 4</p>	<p>$\text{H} \cdot \cdot \ddot{\text{O}} \cdot \cdot \text{H}$</p> <p>$\text{H} \cdot \quad \cdot \ddot{\text{O}} \cdot \cdot \quad \cdot \text{H}$</p>	<p>$\text{H}-\ddot{\text{O}}-\text{H}$</p>

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Lealyn Burr Clapp



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Polar Covalence R Sanderson, 2012-12-02 Polar Covalence provides a detailed account of a successful approach to understanding chemistry from knowledge of atomic structure and the properties that result from this structure This book discusses the nature of multiple bonds Organized into 16 chapters this book begins with an overview of the interrelationships of various basic atomic properties This text then describes chemical bonding which can only occur when the nuclei of both atoms can attract the same electrons Other chapters consider the bond energy of multiple bonds which can be determined by calculating the energy in the usual way as though the bonds were single but of the experimental length This book discusses as well the reduction of the lone pair bond weakening effect through the formation of multiple

bonds The final chapter deals with the relative roles of principles and practice in the teaching of inorganic and general chemistry This book is a valuable resource for chemists and students *The Covalent Bond* Henry Sinclair Pickering,1977

Chemical Binding and Structure J. E. Spice,2016-01-22 Chemical Binding and Structure describes the chemical binding and structure in terms of current chemical theory This book is composed of 13 chapters and starts with a presentation of the principles of the old and modified quantum theory and its application The next chapters cover some basic topics related to chemical binding and structure including electrons the periodic table the electrovalent and covalent bonds and molecular geometry These topics are followed by discussions on the nature of the bond in transition metal complexes electronic and crystal structure crystallinity and other states of matter The concluding chapters are devoted to some analytical techniques for structure determination such as diffraction and spectroscopic methods This book is of value to high school and college chemistry teachers and students Chemical Bonding and the Geometry of Molecules George E. Ryschkewitsch,1963

Structure and Bonding Jack Barrett,2002-05-09 Structure and Bonding covers introductory atomic and molecular theory as given in first and second year undergraduate courses at university level This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes Throughout the book theoretical concepts and experimental evidence are integrated An introductory chapter summarizes the principles on which the Periodic Table is established and describes the periodicity of various atomic properties which are relevant to chemical bonding Symmetry and group theory are introduced to serve as the basis of all molecular orbital treatments of molecules This basis is then applied to a variety of covalent molecules with discussions of bond lengths and angles and hence molecular shapes Extensive comparisons of valence bond theory and VSEPR theory with molecular orbital theory are included Metallic bonding is related to electrical conduction and semi conduction The energetics of ionic bond formation and the transition from ionic to covalent bonding is also covered

Atoms & Chemical Bonding Science Learning Guide NewPath Learning,2014-03-01 The Atoms Atomic Configuration Chemical Bonding Ionic Bonding Ionic Compounds Covalent Bonding Covalent Compounds Naming Compounds and Metallic Bonding Aligned to Next Generation Science Standards NGSS and other state standards Ionic and Covalent Bonding James M. Postma,Julian Roberts,J. Leland Hollenberg,2000-01-15 *Chemistry of the Covalent Bond* Lealyn B. Clapp,1950 Covalent Bonding in Crystals, Molecules, and Polymers James C. Phillips,1969 *Covalent Bonds | Characteristics of Covalent Bonds and Properties of Covalent Compounds | Grade 6-8 Physical Science* Dot EDU,2024-04-15 Dive into the world of chemistry with this essential guide which is perfect for middle schoolers It unravels the complexities of covalent bonds where atoms share electrons to create molecules and how these interactions form diverse substances Ideal for educators homeschooling parents and school librarians this book emphasizes the significance of understanding chemical bonds within the US STEM curriculum Explore covalent compounds fascinating properties and characteristics through engaging explanations and examples This

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