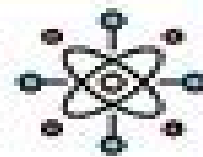


CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CH_4 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions that attract each other. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by a sea of delocalized electrons surrounding positive metal ions. Example: Fe , Cu , Al .

Covalent, Ionic, and Metallic Bonding

Covalent bonding occurs when two atoms share one or more pairs of electrons. This results in a stable molecule. Ionic bonding occurs when one atom transfers an electron to another atom, creating a positive ion and a negative ion. These ions are attracted to each other. Metallic bonding occurs in metals, where the atoms are arranged in a regular lattice, and the outer electrons are delocalized, forming a 'sea of electrons' that holds the metal together.

Science Real

Diagram illustrating the formation of chemical bonds:

The diagram shows the interaction between two atoms, each with a nucleus (positive charge) and a valence shell (outermost shell). The atoms are represented by circles with '+' signs for the nucleus and '-' signs for the valence electrons.

Formation of an Ionic Bond:

One atom (labeled 'A') has a valence shell that is one electron short of being full. The other atom (labeled 'B') has a valence shell that is one electron more than full. Atom A transfers its valence electron to Atom B. This results in Atom A becoming a positive ion (cation) and Atom B becoming a negative ion (anion). The oppositely charged ions are attracted to each other, forming an ionic bond.

Formation of a Covalent Bond:

Two atoms (labeled 'C' and 'D') each have a valence shell that is one electron short of being full. They share their valence electrons, resulting in both atoms having a full valence shell. This sharing of electrons forms a covalent bond.

Formation of a Metallic Bond:

Many metal atoms (labeled 'E') are shown. The outer electrons of these atoms are delocalized, forming a 'sea of electrons' that surrounds the positive metal ions. This sea of electrons holds the metal together, forming a metallic bond.

Diagram illustrating the formation of a covalent bond:

The diagram shows two atoms, each with a nucleus (positive charge) and a valence shell (outermost shell). The atoms are represented by circles with '+' signs for the nucleus and '-' signs for the valence electrons. The atoms share their valence electrons, resulting in a full valence shell for both atoms. This sharing of electrons forms a covalent bond.

Diagram illustrating the formation of an ionic bond:

The diagram shows two atoms, each with a nucleus (positive charge) and a valence shell (outermost shell). The atoms are represented by circles with '+' signs for the nucleus and '-' signs for the valence electrons. One atom transfers its valence electron to the other atom, resulting in a positive ion and a negative ion. The oppositely charged ions are attracted to each other, forming an ionic bond.

Diagram illustrating the formation of a metallic bond:

The diagram shows many metal atoms, each with a nucleus (positive charge) and a valence shell (outermost shell). The atoms are represented by circles with '+' signs for the nucleus and '-' signs for the valence electrons. The outer electrons of these atoms are delocalized, forming a 'sea of electrons' that surrounds the positive metal ions. This sea of electrons holds the metal together, forming a metallic bond.

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Shasha Hu



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