Hess's Law Worksheet

Calculate ΔH for the reaction: C₂H₄ (g) + H₂ (g)→C₂H₅ (g), from the following data.

 $\Delta H = -1411. \text{ kJ}$

$$C_2H_6(g) + 3\% O_2(g) \rightarrow 2 CO_2(g) + 3 H_2O(l)$$

 $\Delta H = -1560$. kJ

$$H_2(g) + \% O_2(g) \rightarrow H_2O(l)$$

 $\Delta H = -285.8 \text{ kJ}$

 Calculate ΔH for the reaction 4 NH₃ (g) + 5 O₂ (g) EB4 NO (g) + 6 H₂O (g), from the following data.

$$N_2(g) + O_2(g) \rightarrow 2 NO(g)$$

 $\Delta H = -180.5 \text{ kJ}$

$$N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$$

 $\Delta H = -91.8 \text{ kJ}$

 $\Delta H = -483.6 \text{ kJ}$

Find ΔH for the reaction 2H₂(g) + 2C(s) + O₂(g) → C₂H₅OH(I), using the following thermochemical data.

 $\Delta H = -875. \text{ kJ}$

$$C(s) + O_2(g) \rightarrow CO_2(g)$$

 $\Delta H = -394.51 \, kJ$

 $\Delta H = -285.8 \text{ kJ}$

Calculate
☐H for the reaction CH₄ (g) + NH₃ (g) → HCN (g) + 3 H₂ (g), given:

$$N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$$

 $\Delta H = -91.8 \text{ kJ}$

 $\Delta H = -74.9 \text{ kJ}$

$$H_2(g) + 2C(s) + N_2(g) \rightarrow 2HCN(g)$$

 $\Delta H = +270.3 \text{ kJ}$

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Peter William Atkins, Julio De Paula

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Table of Contents 12 3 Hess S Law Answers

- 1. Understanding the eBook 12 3 Hess S Law Answers
 - The Rise of Digital Reading 12 3 Hess S Law Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying 12 3 Hess S Law Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 12 3 Hess S Law Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from 12 3 Hess S Law Answers
 - Personalized Recommendations
 - 12 3 Hess S Law Answers User Reviews and Ratings

- 12 3 Hess S Law Answers and Bestseller Lists
- 5. Accessing 12 3 Hess S Law Answers Free and Paid eBooks
 - 12 3 Hess S Law Answers Public Domain eBooks
 - 12 3 Hess S Law Answers eBook Subscription Services
 - 12 3 Hess S Law Answers Budget-Friendly Options
- 6. Navigating 12 3 Hess S Law Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - 12 3 Hess S Law Answers Compatibility with Devices
 - 12 3 Hess S Law Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 12 3 Hess S Law Answers
 - o Highlighting and Note-Taking 12 3 Hess S Law Answers
 - Interactive Elements 12 3 Hess S Law Answers
- 8. Staying Engaged with 12 3 Hess S Law Answers
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 12 3 Hess S Law Answers
- 9. Balancing eBooks and Physical Books 12 3 Hess S Law Answers
 - \circ Benefits of a Digital Library
 - Creating a Diverse Reading Collection 12 3 Hess S Law Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 12 3 Hess S Law Answers
 - Setting Reading Goals 12 3 Hess S Law Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 12 3 Hess S Law Answers
 - Fact-Checking eBook Content of 12 3 Hess S Law Answers
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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