

Once this is done, the mass of the gas can be computed from the amount and the molar mass:

$$m = (1.66 \times 10^2 \text{ mol}) \times (16.04 \text{ g mol}^{-1}) = 2.67 \times 10^3 \text{ g} = \boxed{2.67 \times 10^3 \text{ kg}}$$

**E1.5(b)** Use the equation for hydrostatic pressure [1.3],  $p = p_{\text{ex}} + \rho gh$ . Let  $p_{\text{ex}}$  be the pressure at the top of the straw and  $p$  is the pressure on the surface of the liquid (atmospheric pressure). Thus, the pressure difference is

$$\begin{aligned} p - p_{\text{ex}} &= \rho gh = (1.0 \text{ g cm}^{-3}) \times \frac{1 \text{ kg}}{10^3 \text{ g}} \times \left( \frac{1 \text{ cm}}{10^{-2} \text{ m}} \right)^3 \times (9.81 \text{ m s}^{-2}) \times (0.15 \text{ m}) \\ &= \boxed{1.5 \times 10^3 \text{ Pa}} = 1.5 \times 10^{-2} \text{ atm} \end{aligned}$$

**E1.6(b)** The pressure in the apparatus is given by

$$p = p_{\text{ex}} + \rho gh \quad [1.3]$$

where  $p_{\text{ex}} = 760 \text{ Torr} = 1 \text{ atm} = 1.013 \times 10^5 \text{ Pa}$ ,

and  $\rho gh = 13.55 \text{ g cm}^{-3} \times \left( \frac{1 \text{ kg}}{10^3 \text{ g}} \right) \times \left( \frac{1 \text{ cm}}{10^{-2} \text{ m}} \right)^3 \times 0.100 \text{ m} \times 9.806 \text{ m s}^{-2} = 1.33 \times 10^4 \text{ Pa}$

$$p = 1.013 \times 10^5 \text{ Pa} + 1.33 \times 10^4 \text{ Pa} = 1.146 \times 10^5 \text{ Pa} = \boxed{115 \text{ kPa}}$$

**E1.7(b)** Rearrange the perfect gas equation [1.8] to give  $R = \frac{pV}{nT} = \frac{pV_{\text{m}}}{T}$

All gases are perfect in the limit of zero pressure. Therefore, the value of  $pV_{\text{m}}/T$  extrapolated to zero pressure will give the best value of  $R$ .

The molar mass can be introduced through

$$pV = nRT = \frac{m}{M} RT$$

which on rearrangement gives  $M = \frac{m}{V} \frac{RT}{p} = \rho \frac{RT}{p}$ .

The best value of  $M$  is obtained from an extrapolation of  $\rho/p$  versus  $p$  to zero pressure; the intercept is  $M/RT$ .

Draw up the following table:

$p/\text{atm}$	$(pV_{\text{m}}/T)/(\text{dm}^3 \text{ atm K}^{-1} \text{ mol}^{-1})$	$(\rho/p)/(\text{g dm}^{-3} \text{ atm}^{-1})$
0.750 000	0.082 0014	1.428 59
0.500 000	0.082 0227	1.428 22
0.250 000	0.082 0414	1.427 90

From Figure 1.1(a),  $\left( \frac{pV_{\text{m}}}{T} \right)_{p \rightarrow 0} = \boxed{0.082 062 \text{ dm}^3 \text{ atm K}^{-1} \text{ mol}^{-1}}$

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