

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from  $\text{pH}$  and  $\text{pOH}$ , and vice versa from  $\text{pH}$  and  $\text{pOH}$ .
3. Calculate the  $\text{pH}$  of a weak acid.

## Vocabulary

- |                   |                               |                  |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. $\text{pH}$                | 6. $\text{pOH}$  |

## Key Equation

- |                                    |                                  |
|------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

## Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the problem. Express answers in scientific notation when appropriate.

1. What substance has a  $\text{pH}$  of 10.0? (hydrogen ion concentration)
2. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
3. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
4. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
5. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
6. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
7. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
8. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
9. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
10. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
11. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
12. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
13. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
14. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
15. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
16. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
17. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
18. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
19. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)
20. What substance has a  $\text{pH}$  of 1.0? (hydrogen ion concentration)

1.  $1.0 \times 10^{-14}$
2.  $1.0 \times 10^{-14}$
3.  $1.0 \times 10^{-14}$
4.  $1.0 \times 10^{-14}$
5.  $1.0 \times 10^{-14}$
6.  $1.0 \times 10^{-14}$
7.  $1.0 \times 10^{-14}$
8.  $1.0 \times 10^{-14}$
9.  $1.0 \times 10^{-14}$
10.  $1.0 \times 10^{-14}$
11.  $1.0 \times 10^{-14}$
12.  $1.0 \times 10^{-14}$
13.  $1.0 \times 10^{-14}$
14.  $1.0 \times 10^{-14}$
15.  $1.0 \times 10^{-14}$
16.  $1.0 \times 10^{-14}$
17.  $1.0 \times 10^{-14}$
18.  $1.0 \times 10^{-14}$
19.  $1.0 \times 10^{-14}$
20.  $1.0 \times 10^{-14}$

# 19 2 Hydrogen Ions And Acidity Worksheet Answers

**Sabine Zange**



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