## Hess's Law Worksheet

Calculate ΔH for the reaction: C<sub>2</sub>H<sub>4</sub> (g) + H<sub>2</sub> (g)→C<sub>2</sub>H<sub>5</sub> (g), from the following data.

 $\Delta H = -1411. \text{ kJ}$ 

$$C_2H_6(g) + 3\% O_2(g) \rightarrow 2 CO_2(g) + 3 H_2O(l)$$

 $\Delta H = -1560$ . kJ

$$H_2(g) + \% O_2(g) \rightarrow H_2O(l)$$

 $\Delta H = -285.8 \text{ kJ}$ 

 Calculate ΔH for the reaction 4 NH<sub>3</sub> (g) + 5 O<sub>2</sub> (g) EE4 NO (g) + 6 H<sub>2</sub>O (g), from the following data.

$$N_2(g) + O_2(g) \rightarrow 2 NO(g)$$

 $\Delta H = -180.5 \text{ kJ}$ 

$$N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$$

 $\Delta H = -91.8 \text{ kJ}$ 

 $\Delta H = -483.6 \text{ kJ}$ 

Find ΔH for the reaction 2H<sub>2</sub>(g) + 2C(s) + O<sub>2</sub>(g) → C<sub>2</sub>H<sub>5</sub>OH(I), using the following thermochemical data.

 $\Delta H = -875. \text{ kJ}$ 

$$C(s) + O_2(g) \rightarrow CO_2(g)$$

 $\Delta H = -394.51 \, kJ$ 

 $\Delta H = -285.8 \text{ kJ}$ 

Calculate 
☐H for the reaction CH<sub>4</sub> (g) + NH<sub>3</sub> (g) → HCN (g) + 3 H<sub>2</sub> (g), given:

$$N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$$

 $\Delta H = -91.8 \text{ kJ}$ 

 $\Delta H = -74.9 \text{ kJ}$ 

$$H_2(g) + 2C(s) + N_2(g) \rightarrow 2HCN(g)$$

 $\Delta H = +270.3 \text{ kJ}$ 

# 12 3 Hess S Law Answers

**Scott C. Dulebohn** 

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#### **Table of Contents 12 3 Hess S Law Answers**

- 1. Understanding the eBook 12 3 Hess S Law Answers
  - The Rise of Digital Reading 12 3 Hess S Law Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying 12 3 Hess S Law Answers
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - $\circ \ \ Determining \ Your \ Reading \ Goals$
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 12 3 Hess S Law Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from 12 3 Hess S Law Answers
  - Personalized Recommendations

- 12 3 Hess S Law Answers User Reviews and Ratings
- 12 3 Hess S Law Answers and Bestseller Lists
- 5. Accessing 12 3 Hess S Law Answers Free and Paid eBooks
  - 12 3 Hess S Law Answers Public Domain eBooks
  - 12 3 Hess S Law Answers eBook Subscription Services
  - 12 3 Hess S Law Answers Budget-Friendly Options
- 6. Navigating 12 3 Hess S Law Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - 12 3 Hess S Law Answers Compatibility with Devices
  - 12 3 Hess S Law Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 12 3 Hess S Law Answers
  - Highlighting and Note-Taking 12 3 Hess S Law Answers
  - Interactive Elements 12 3 Hess S Law Answers
- 8. Staying Engaged with 12 3 Hess S Law Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - o Following Authors and Publishers 12 3 Hess S Law Answers
- 9. Balancing eBooks and Physical Books 12 3 Hess S Law Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 12 3 Hess S Law Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 12 3 Hess S Law Answers
  - Setting Reading Goals 12 3 Hess S Law Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 12 3 Hess S Law Answers
  - Fact-Checking eBook Content of 12 3 Hess S Law Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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