

- (d) A sodium hydroxide solution was made up by dissolving 20.0 g of sodium hydroxide in water and making the volume of the solution up to 1.00 dm³. Calculate the concentration of sodium hydroxide, NaOH, in this solution in mol dm⁻³.

(relative atomic masses: H = 1.00, O = 16.0, Na = 23.0)

$$1 + 16 + 23 = 40 \quad \frac{V}{M} = C \quad (2)$$

$$\frac{20}{40} = 0.5$$

concentration = 0.5 mol dm⁻³

- (e) In another experiment, a titration was carried out. 25.0 cm³ of 1.50 mol dm⁻³ sodium hydroxide solution, NaOH, was titrated with hydrochloric acid. The volume of the hydrochloric acid required to neutralise the sodium hydroxide solution was 30.0 cm³.

Calculate the concentration of the hydrochloric acid, HCl, in mol dm⁻³.

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Jin-Ying Zhang



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