

Name _____

Date ____ / ____ / ____

ATOMIC NUMBERS, PROTONS, NEUTRONS AND ELECTRONS

The number of protons in each atom of an element is its atomic number. The atomic number is also the number of electrons since atoms themselves are electronically neutral. Because neutrons and protons are almost the same mass, the total number of protons and neutrons in an atom is the atomic mass. Therefore, by subtracting the atomic number from the atomic mass you can calculate the number of neutrons.

This is why the mass number and atomic number, but not the number of neutrons, is shown in the periodic table. You can use the information provided below to determine the missing values in the chart below.



Example: Helium atom



Use this knowledge and the information provided below to determine the missing values in the chart below.

Element	Symbol	Atomic Number	Atomic Mass	Protons	Neutrons	Electrons
Hydrogen	H	1	1	1	0	1
Helium	He	2	4	2	2	2
Lithium	Li	3	7	3	4	3
Boron	B	5	11	5	6	5
Carbon	C	6	12	6	6	6
Nitrogen	N	7	14	7	7	7
Oxygen	O	8	16	8	8	8
Fluorine	F	9	19	9	10	9
Neon	Ne	10	20	10	10	10
Sodium	Na	11	23	11	12	11
Magnesium	Mg	12	24	12	12	12
Aluminum	Al	13	27	13	14	13
Silicon	Si	14	28	14	14	14
Phosphorus	P	15	31	15	16	15
Sulfur	S	16	32	16	16	16
Chlorine	Cl	17	35	17	18	17
Argon	Ar	18	40	18	22	18
Potassium	K	19	39	19	20	19
Calcium	Ca	20	40	20	20	20

Where Are The Electrons Answer Key

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Table of Contents Where Are The Electrons Answer Key

1. Understanding the eBook Where Are The Electrons Answer Key
 - The Rise of Digital Reading Where Are The Electrons Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Where Are The Electrons Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Where Are The Electrons Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Where Are The Electrons Answer Key
 - Personalized Recommendations
 - Where Are The Electrons Answer Key User Reviews and Ratings

- Where Are The Electrons Answer Key and Bestseller Lists
- 5. Accessing Where Are The Electrons Answer Key Free and Paid eBooks
 - Where Are The Electrons Answer Key Public Domain eBooks
 - Where Are The Electrons Answer Key eBook Subscription Services
 - Where Are The Electrons Answer Key Budget-Friendly Options
- 6. Navigating Where Are The Electrons Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Where Are The Electrons Answer Key Compatibility with Devices
 - Where Are The Electrons Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Where Are The Electrons Answer Key
 - Highlighting and Note-Taking Where Are The Electrons Answer Key
 - Interactive Elements Where Are The Electrons Answer Key
- 8. Staying Engaged with Where Are The Electrons Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Where Are The Electrons Answer Key
- 9. Balancing eBooks and Physical Books Where Are The Electrons Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Where Are The Electrons Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Where Are The Electrons Answer Key
 - Setting Reading Goals Where Are The Electrons Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Where Are The Electrons Answer Key
 - Fact-Checking eBook Content of Where Are The Electrons Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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