

Chapter 5 The Periodic Table

Section 5.2 The Modern Periodic Table

(pages 136–138)

This section explains the organization of the modern periodic table and discusses the general properties of metals, nonmetals, and metalloids.

Reading Strategy (page 138)

Precirvng: Before you read, complete the table by writing two questions about the periodic table on pages 132–133. As you read, write answers to your questions. For more information on this reading strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Questions About the Periodic Table	
Description	Answer
Quesion: What element was first discovered on the basis of its atomic mass?	Atomic mass is a value that describes an element's mass based on the mass of carbon-12.
Question: Why does the position of an element predict its behavior in other elements?	The number of valence electrons determines how an element reacts with other elements.

Other possible questions: "Why are there less nonmetals than metals in period 2?" "Why is period 7 incomplete?"

The Periodic Law (pages 134–135)

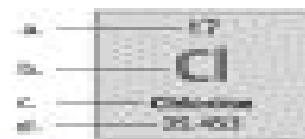
1. Is the following sentence true or false? In the modern periodic table, elements are arranged by increasing number of protons.

True

2. Explain why the number of elements per period varies. The number of electrons in a given period varies because the number of available orbits increases from one to the next.
3. Properties of elements repeat in a predictable way when atomic numbers are used to arrange elements into groups. This pattern of repeating properties is called the periodic law.

Atomic Mass (page 134)

4. Label the four types of information supplied for chlorine in the diagram.



a. Atomic number

b. Element symbol

c. Element name

d. Atomic mass

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