Types of Bonds

Look at the Bohr Models below, move the eletrons from one model to the other as they would if these two atoms formed a bond.

A.

After moving the electrons.

How many Protons * How many Electrons * How many Electrons * How many Electrons * How many Electrons * What is the difference * How many Electrons * How

1	16	sufficiency.	A	Louises	electrons	BOX.	a divorce	B.
	18.00	THE RESERVE OF THE RE	40.00	\$100 Dec 100	ACTUAL VALUE OF PERSONS IN CO.	100 100 100	ACRES OF ACTUAL PROPERTY.	B-17 (F)

- a. how many electrons will atom A lose? ______
- b. how many electrons will atom B gain? ______
- c. what will be the oxidation number of atom A? (include sign)
- d. what will be the oxidation number of atom B? [Include sign]
- e. what will be the total charge of the compound formed? Number format.
- f. what type of bond will form?
- Explain why an element's oxidation number is related to the group on the periodic table to which it belongs.

For Families 1, 2, 3, and 4 the oxidation number is

For Families 5,6,7, and 8 the oxidation number is

Directions: Complete the table comparing ionic compounds and covalent compounds.

Characteristic	lonic compounds	Covalent compounds		
3. How the compound is formed	electrons are shared electrons are transferred	electrons are shared electrons are transferred		
4. Smallest particle	proton, neutron, electron	proton, neutron, electron		
Usual state at room temperature	Solid, Liquid, Gas	Solid, Liquid, Gas		

Where Are Ionic Bonds Worksheet 233 Physical Science Answers

Julia Schneider

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