

Tests for functional groups

• Alkenes (C=C)

i) Reaction with bromine water (electrophilic addition)

Shake 1 cm³ of the compound with 1 cm³ of bromine water (or bubble it through if a gas) in the dark. If the orange bromine is decolourised, this confirms an alkene



ii) Reaction with acidified potassium manganate(VII) (oxidative addition)

Shake 1 cm³ of the compound with 1 cm³ of potassium manganate(VII) solution acidified with a little dilute sulphuric acid (or bubble it through if a gas). If the purple manganate(VII) is decolourised, this confirms an alkene



• Halogenoalkanes (R-X)

Reaction with aqueous sodium hydroxide (nucleophilic substitution)

Add 1 cm³ of the compound to 1 cm³ of aqueous sodium hydroxide in a test tube and heat to 50°C in a water bath; neutralise excess hydroxide with dilute nitric acid; add 1 cm³ of aqueous silver nitrate. White ppt of AgCl \Rightarrow R-Cl; cream ppt of AgBr \Rightarrow R-Br; yellow ppt of AgI \Rightarrow R-I



• Alcohols (R-OH)

Reaction with phosphorus pentachloride

Add a spatula measure of PCl₅ solid to 1 cm³ of the liquid to be tested. An exothermic reaction which evolves steamy fumes (HCl) which turn damp litmus red and form a white 'smoke' with ammonia gas confirms the presence of an -OH group in the molecule. As this also gives a positive result with water (H-OH) and carboxylic acids (R-COOH), you must also be sure that the 'unknown' is dry and neutral before you can confirm it is an alcohol



Once you have confirmed an alcohol is present, you may need to identify it as primary, secondary or tertiary

a) Primary alcohols (R-CH₂-OH)

Oxidise with potassium dichromate acidified with a little dilute sulphuric acid, distilling off the product. If the product shows a positive test for an aldehyde (see below) the 'unknown' must be a primary alcohol

b) Secondary alcohols (R₂CH-OH)

Oxidise with potassium dichromate acidified with a little dilute sulphuric acid, distilling off the product. If the product shows a positive test for a carbonyl group but a negative test for an aldehyde (see below) the product must be a ketone and the 'unknown' must be a secondary alcohol

c) Tertiary alcohols (R₃C-OH)

Attempt to oxidise by warming with potassium dichromate acidified with a little dilute sulphuric acid. If the potassium dichromate remains orange, no oxidation has taken place and the 'unknown' must be a tertiary alcohol

d) Compounds containing -CH(CH₃)OH – the iodoform (triiodomethane) reaction

To 2 cm³ of aqueous sodium hydroxide add 1 cm³ of the 'unknown' followed by a few drops of iodine solution, stopper and shake well. The production of a pale yellow crystalline precipitate (of triiodomethane) confirms the presence of either the -CH(CH₃)OH group (present in secondary methyl alcohols or ethanol) or a methyl ketone (see below) which must be distinguished by additional tests.

• Carbonyl compounds (R₂C=O)

Reaction with 2,4-dinitrophenylhydrazine (condensation)

Add a few drops of the 'unknown' to 1 cm³ of a solution of 2,4-DNP in methanol. A dense yellow/orange precipitate confirms the presence of a carbonyl group



Once a carbonyl group has been identified, it may be confirmed as either an aldehyde or ketone by one of the following procedures:

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Raymond J. Lukens



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The Hidden Curriculum—Faculty-Made Tests in Science Sheila Tobias, Jacqueline Raphael, 2013-06-29 This resource manual for college level science instructors reevaluates the role of testing in their curricula and describes innovative techniques pioneered by other teachers part I examines the effects of the following on lower division courses changes in exam content format and environment revisions in grading practices student response colleague reaction the sharing of new practices with other interested professionals and more The book includes a comprehensive introduction faculty composed narratives commentaries by well known science educators and a visual index to 100 more refined innovations *Peterson's Graduate Programs in the Physical Sciences, Mathematics, Agricultural Sciences, the Environment & Natural Resources* 2012 Peterson's, 2011-12-30 Graduate Programs in the Physical Sciences Mathematics Agricultural Sciences the Environment Natural Resources 2012 contains more than 2 900 graduate programs in 59 disciplines including agriculture and food sciences astronomy and astrophysics chemistry physics mathematics environmental sciences and management natural resources marine sciences and more This guide is part of Peterson's six volume Annual Guides to Graduate Study the only annually updated reference work of its kind provides wide ranging information on the graduate and professional programs offered by U S accredited colleges and universities in the United States and throughout the world Informative data profiles for more than 2 900 graduate programs in 59 disciplines including facts and figures on accreditation degree requirements application deadlines and contact information financial support faculty and student body profiles Two page in depth descriptions written by featured institutions offer complete details on specific graduate programs schools or departments as well as information on faculty research and the college or university Expert advice on the admissions process financial support and accrediting agencies Comprehensive directories list programs in this volume as well as others in the graduate series Up to date appendixes list institutional changes since the last addition along with abbreviations used in the guide

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