Answers

AS Level

Topic 1: Chemical formulae and moles

Answers to in-text questions

Questions 1.1

- 1 1.5 mol
- 2 0:75 mol
- 3 0:667 mol

Questions 1.2

- 1 16.7 mol
- 2 8.3×10⁻³ mol

Questions 1.3

- 3 5
- 2 11
- 3 (2
- 4 24
- 5 22

Questions 1.4

- 1 a) 151.9
 - b) 162.1
 - c) 60.0
 - d) 132.1 e) 388.8
- 2. a) 0.5 mol
 - b) 2.5 mol
 - b) 2.5 molc) 0.6 mol
 - d) 2.0 mol
 - d) 2.0 mol
 e) 0.38 mol
- 3 a) 180.6g
- b) 44.46g

Questions 1.5

- 1 M, = 277.9 moles of oxygen = 11
- M_s = 474.3 moles of exygen = 20

Questions 1.6

- 1 Cu₂5
- 2 C₃H₆
- 3 CaFe₂O₄

Questions 1.7

- 1 a) 2H₂ + O₂ → 2H₂O
 - b) l₂ + 3Cl₂ → 21Cl₃
 - c) NaOH + Al(OH)3 -> NaAlO3 + 2H2O
 - d) 2H₂S + SO₂ → 3S + 2H₂O
 - e) 4NH₃ + 3O₂ → 2N₂ + 6H₂O

- 2 a) MgCO₅ → MgO + CO₂
 - b) $Pb(s) + 2AgNO_3(aq) \rightarrow Pb(NO_3)_3(aq) + 2Ag(s)$
 - c) $Na_2O(s) + H_2O(l) \rightarrow 2NaOH(aq)$
 - d) 2FeO₂ + Cl₂ → 2FeO₃
 - E) Fe₂(SO₄)₃ + 6NaOH → 2Fe(OH)₃ + 3Na₂SO₄

Questions 1.8

- 1 170g
- 2 60.7g
- 3 51.1%

Questions 1.9

- 1 1.2 dm²
- 2 Volume of 5 O₂ = 50 cm³, mass of sulfur = 0.20 g

Questions 1.10

- 1 a) 2.67 moldm⁻³
 - b) 0.025 mol dm⁻³
 - c) 0.833 mol dm⁻³
- 2 a) 0.75 mol
 - b) 0.044 mol
 - a) 0.010 mol

Questions 1.11

- 1 1.2 dm³
- 2 12 dm³
- 3 Volume of acid = 0.41 dm³; volume of H₂ = 4.9 dm³
- 4 0.64g
- 5 1.27 moldm⁻³

Answers to end of topic questions

- 1 a) Any two of ZnCO₃, Zn(OH)₂ and ZnO
 - To make sure no water of crystallisation remained
 - iii) Mass of zinc sulface = mass of tube and salt after heating - mass of tube = 76.34 - 74.25 = 2.09g Relative molecular mass of zinc sulface (ZnSO₄) = 65.4 + 32.1 + (4 × 16.0) = 161.5 Amount in moles = 2.09 + 161.5 = 0.01294 = 1.29 × 10⁻⁴
 - (iii) Mass of water = mass of tube and hydrated salt = mass of tube and salt after heating = 77.97 76.34 = 1.63g. Molar mass of water (H₂O) = (2 × 1) + 16 = 18. Amount in moles = 1.63 + 18 = 0.0905 = 9.1 × 10⁻².
 - iv) Value of $x = 9.1 \times 10^{-2} + 1.29 \times 10^{-2} = 7.054$ which is 7 as a whole number

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