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web learners have completed the relevant section of the specification need to know that bases are chemicals which react with acids in neutralisation reactions they are usually metal oxides metal hydroxides or metal carbonates examples are copper oxide and copper carbonate many bases are insoluble

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web an acid is any hydrogen containing substance that is capable of donating a proton hydrogen ion to another substance a base is a molecule or ion able to accept a hydrogen ion from an acid acidic substances are usually identified by their sour taste

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combined the acid is neutralized by the base the products of this reaction are an ionic compound which is labeled as a salt and water

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logarithmic scale of relative h concentration

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chemistries the acid and base cancel each other's chemistry to produce a rather innocuous substance water in fact the general reaction between an acid and a base is  $\text{acid} + \text{base} \rightarrow \text{water} + \text{salt}$

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