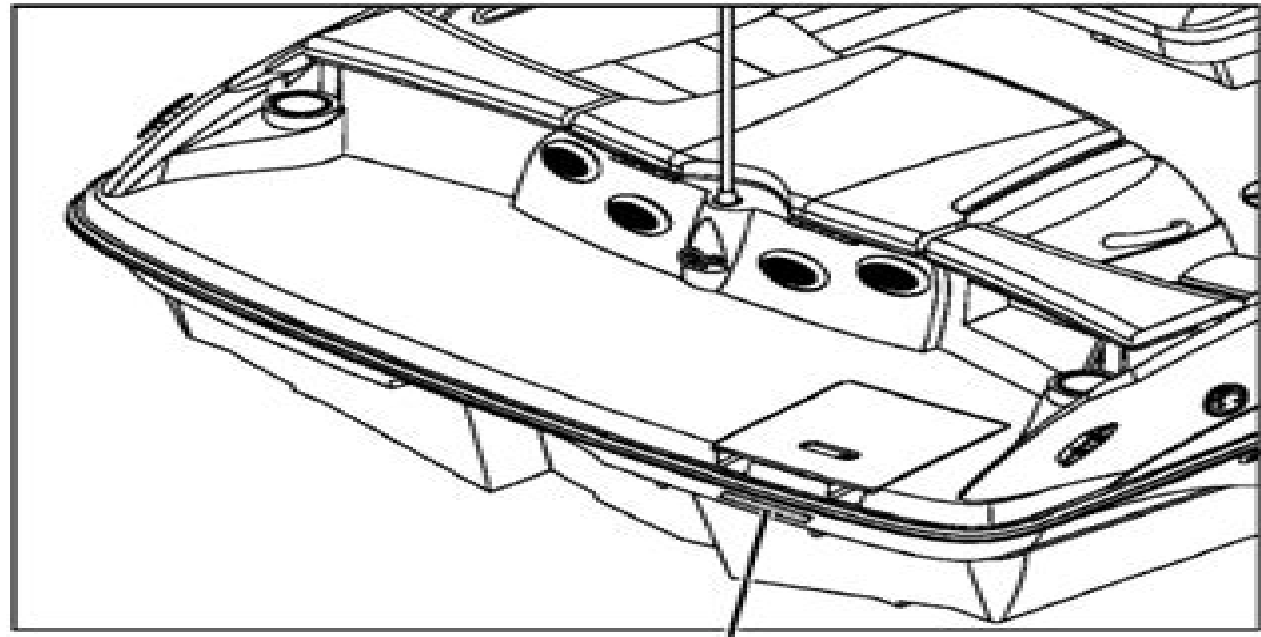




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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times \frac{101.96 \text{ g Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} = 81.18 \text{ g Al}_2\text{O}_3$ c determine the percent yield $\frac{66.0 \text{ g}}{81.18 \text{ g}} \times 100 = 81.3\%$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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