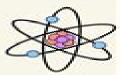


> ATOMIC STRUCTURE 2



Attorne structure -

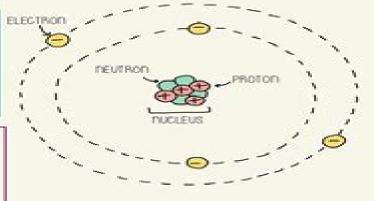
Every oform is made up of a central nucleus surrounded by electrons whitting around the nucleus in electron arbitals. The nucleus is they in relation to the size of the whole atom - comparable to a pea in the middle of a faotbal stadium. The nucleus is made up of profate and neutrons, which are the heavest particles. Bestrons are approximately 2000 times lighter than a proton or restron and weigh procladly nothing, therefore most of the mass comes from the tiny nucleus Profate are post-vely sharged whereas resultants, have no charge, making the nucleus positively charged overall Electrons are negatively charged, which enables the nucleus to keep haid of its electrons, due to the attraction between the apposite charges.

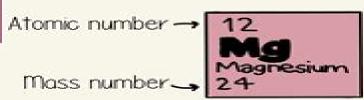
- Allomo number & mass number -

Each atom has an atomic number and a mass number, which are displayed on the Periodic Table as the small number and larger number next to each chemical symbol. The latomic number lets us the number of protons which is unique for each element. If the number of protons changes land therefore the atomic number shanges), which could occur during radioactive decay, for example, then the identity of the element also changes.

The mass number stells us the number of both protons and neutrons in the nucleus. Remember that most of an atom's mass comes from the nucleus and that the nucleus is made up of both protons and neutrons. Two atoms of the same element can have different mass numbers because they are isotopes.

Atomic particle	Charge	Mass
Proton	-1	
Deutron	0	
Electron		1/2000





- Isotopes -

befores are aforms which have the same number of profined but different numbers of newhold. Therefore they have the same atoms, number but a different mass number. For example, the most common safepes of carbon are carbon-12 and carbon-13, in which carbon-13 has an additional newtron in its nucleus to give it a higher mass number.

The mass numbers that we see in the periodic table are an average of all the masses of all satapes of a particular element. We call this the install we allow a formula and can be calculated using the following formula:

(Mass number of isotope 1 x percentage of isotope 1) • (mass number of isotope 2 x percentage of isotope 2)

Worked example:

Lets say we have a sample of carbon which consists of 76% of the solope carbon-12 and 24% carbon-13. Then the relative atomic mass (RATI) would be

RAM - (76 x 12) - (24 x 13) 100

BATS - 1229

Overshoos

- I flame the two subolamic particles that are found in the nucleus and state their charges
- 2. What does atomic number represent?
- 3. Why do atoms have an overall neutral charge?
- 4 Sadam Dalihas on atomic number of 11 and a mass number of 23 State. The number or profoss, neutrons and electrons in a sadam atom.
- 5 What is an isotope?"
- 6. A sample of bramine contains 45 / bramine-79 and 55 / bramine-81. Cabulate its relative atomic mass (RAMI)

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S Baum

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