

## Free Energy

When does the energy come from the change of things being done?

### What

All things change because of changes in energy. This is not surprising. The only way to change the amount of energy in an environment is to take an amount of energy out of it, or add it. Thus, all the things that are surrounded with us, that exist in part of the layer of conservation of energy, in the world, that energy for energy energy must come.

### Model 1 – Spontaneous Processes

|   | Process Description                                                              | Change in Enthalpy (ΔH) | Change in Entropy (ΔS) | Spontaneous? |
|---|----------------------------------------------------------------------------------|-------------------------|------------------------|--------------|
| a | Hot water (100°C) → Cold water (20°C)                                            | –                       | Increasing             | Yes          |
| b | Hot water (100°C) → Cold water (20°C) + Hot water (100°C) → Cold water (20°C)    | –                       | Increasing             | Yes          |
| c | Hot water (100°C) → Cold water (20°C) + Hot water (100°C) → Cold water (20°C)    | Decreasing (negative)   | Decreasing (negative)  | Yes          |
| d | $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{s})$ at 100°C | Decreasing              | Decreasing             | Yes          |
| e | $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{s})$ at 20°C  | Decreasing              | Decreasing             | Yes          |
| f | $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{s})$          | Decreasing              | Decreasing             | Yes          |
| g | $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{s})$          | Decreasing              | Decreasing             | Yes          |
| h | $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{s})$          | Decreasing              | Decreasing             | Yes          |

#### 1. Answer Model 1:

- When the process is spontaneous, the change in enthalpy (ΔH) is negative.
- When the process is spontaneous, the change in entropy (ΔS) is positive.
- When the process is spontaneous, the change in enthalpy (ΔH) is negative.
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- When the process is spontaneous, the change in enthalpy (ΔH) is negative.

# Atp The Free Energy Carrier Pogil Activity Answers

**Huangqi Zhang**



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