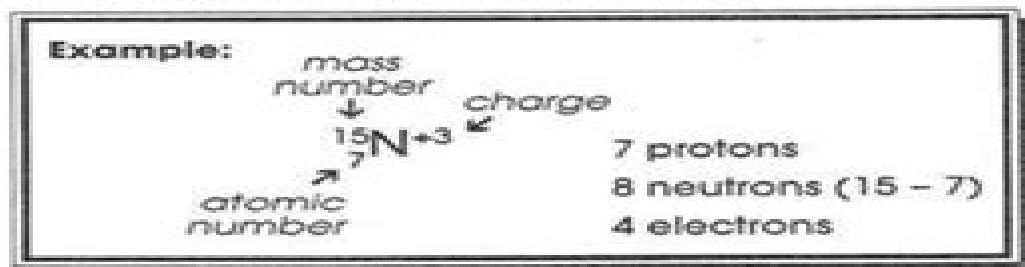


ATOMIC STRUCTURE

Name _____

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.



Complete the following chart.

Element/ Ion	Atomic Number	Atomic Mass	Mass Number	Protons	Neutrons	Electrons
H						
H ⁺						
$^{12}_6\text{C}$						
$^7_3\text{Li}^+$						
$^{35}_{17}\text{Cl}^-$						
$^{39}_{19}\text{K}$						
$^{24}_{12}\text{Mg}^{2+}$						
As ³⁻						
Ag						
Ag ⁺¹						
S ⁻²						
U						

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Atomic and Molecular Structure Joel Chaffee, 2011-01-15 Learn about the atom what it is the people responsible for helping us understand it and how it affects us in the world today **The Basics of Atoms and Molecules** Tom Jackson, 2013-12-15 A fun filled introduction to matter the elements of the periodic table atoms electrons reactions and bonding and radioactivity this volume provides young adults with chemistry examples that reflect their real world interconnections in science Key terms easy experiments and clear illustrations help to guide students through chemical applications A chapter about Niels Bohr and his model for the atom honors his contribution to the understanding of atomic structure and to nuclear fission Tools and techniques such as a scanning tunneling microscope Rutherford's gold foil experiment and a mass spectrometer highlight this instructive text that is aligned to the Common Core Standards

Chemistry & Atomic Structure John David Main Smith, 1924 **Chemistry for Kids, the Atomic Structure** Lex Sharp, 2018-03-24 This book is meant for students grades 6 and higher It addresses concepts that are critical to the understanding of the inner structure of the atom which lays the foundation for appreciating how elements behave in chemical reactions The text has been integrated with colorful images that help engage the visual with the conceptual No prerequisites are necessary The book is suitable for the beginner who has not studied chemistry before and for the student who wants to clarify questions for review purposes A couple of mathematical chapters were also infused into the book to support younger learners with the necessary mathematical background to succeed To get the most from this book students should summarize each chapter and rehearse the concepts Attempting to work independently through all problems at the end of the book is beneficial only before reading the solutions When solving problems one can be sure the chemical principles presented were well understood Restricting oneself to only reading is incomplete learning The series walks learners gradually and conceptually through the chemical and particle physics theories that are meant to jumpstart a more profound scientific understanding Topics covered in this volume Subatomic Particle Atomic Mass Period Table of Elements Atomic Number Isotope and Ion Notations Atomic Weight The Mole Molar Mass Exponents The Rule of Three Practice problems and easy to follow solutions **Atomic Structure and Valency**, 1969 **ATOMIC STRUCTURE High School Chemistry**

Curriculum Bud Morgan, 2025-07-09 Welcome to our comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry In this book we will explore the fascinating world of atoms the basic building blocks of matter and delve into the complexities of their structure Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the development of modern chemistry We will examine the key scientists who contributed to our understanding of atomic structure from Democritus to Thomson Rutherford and Bohr Moving on to Chapter 2 we will explore the structure of the atom in more detail including the subatomic particles that make up an atom protons neutrons and electrons We will also discuss the concept of atomic number atomic mass and isotopes In Chapter 3 we

will delve into the arrangement of electrons within an atom's electron cloud including energy levels, sublevels, and orbital shapes. We will also discuss the principles of electron configuration and how it impacts an atom's chemical properties. Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure. We will explore periodic trends such as atomic radius, ionization energy, and electron affinity, and discuss how these trends are related to an element's position on the table. In our final chapter, we will discuss the significance of atomic structure in various chemical reactions, including bonding, stoichiometry, and the behavior of gases. We will also explore the importance of understanding atomic structure in fields such as biochemistry, environmental science, and materials science. Throughout this book, we will provide clear explanations, helpful illustrations, and interactive activities to reinforce key concepts. Whether you are a visual learner, a hands-on learner, or a theoretical learner, this book is designed to meet your needs. We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry. So let's dive in and unlock the mysteries of the atom together.

Elementary Atomic Structure Gordon Kemble Woodgate, 1998
Chemistry and Atomic Structure Jd Main Smith, 2023-07-18 This book is an introduction to the chemistry of atomic structure. It covers the basic principles of atomic theory, including the nature of atoms, chemical bonds, and chemical reactions. The author provides clear explanations of complex concepts, making this book accessible to students and non-experts alike. Anyone interested in chemistry or atomic theory will find this book a useful resource. This work has been selected by scholars as being culturally important and is part of the knowledge base of civilization as we know it. This work is in the public domain in the United States of America and possibly other nations. Within the United States, you may freely copy and distribute this work as no entity individual or corporate has a copyright on the body of the work. Scholars believe and we concur that this work is important enough to be preserved, reproduced, and made generally available to the public. We appreciate your support of the preservation process and thank you for being an important part of keeping this knowledge alive and relevant.

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