

Name: _____

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S'mores Stoichiometry Lab



Introduction:

In this experiment, you will explore the principles of stoichiometry by building S'mores, the delicious, chocolate, marshmallow, and graham cracker treats.

Definitions: 'Stoichio' means element and 'metry' means the process of measuring. The mass and quantity relationships among reactants and products in a reaction are found using the process of stoichiometry.

Purpose:

To determine the limiting reactant in the synthesis of S'mores.

Prelab Questions:

1. What does a chemical equation tell us?
2. What do the coefficients in a balanced chemical equation represent?

Materials:

Miniature Chocolate bars
Marshmallows
Bunsen Burner

Graham cracker Square
Paper Towels
Wooden skewers

Reference Data:

Substance	Symbol	Unit mass	Package Mass
Graham Cracker Squares	Gr	8.50 g	408 g
Marshmallow	Mm	7.125 g	285 g
Miniature Chocolate Bar	Ch	12.00 g	348 g

** for the purposes of this lab the UNIT mass and MOLAR mass are the same thing

Procedure:

1. Perform a synthesis reaction (forming ~~one~~ s'more). For this lab, we will consider a s'more to consist of 2 graham cracker squares, 1 marshmallow and 1 miniature chocolate bar.
 - a) Determine the formula for a s'more
 - b) write the balanced chemical equation using the symbols above
 - c) According to the balanced equation complete the following ratios:

_____ Gr = _____ Mm	_____ Gr = _____ S'more
_____ Ch = _____ Mm	_____ Mm = _____ S'more
_____ Ch = _____ Gr	_____ Ch = _____ S'more
2. Calculating the unit mass of the s'more.
 - a) Calculate the mass on ONE s'more.
 - b) Is there a relationship between the mass of a S'more and the masses of the reactants used to make it?
 - c) If so, what is the relationship?
 - d) What law have you studied in this course that might define this relationship?
3. Predicting the limiting reactant:
 - a) Complete the data table below based on the materials provided for your group.

Substance	Quantity received	# of S'mores it COULD make
Graham Cracker Squares		
Marshmallow		
Miniature Chocolate Bar		
 - b) What will the limiting reactant be?
 - c) How many of the other substances will remain
 - d) What are the masses of each excess reactant that will remain once the reaction is complete?

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