

Topic Name : Atomic Structure

1. The mass of a proton is equal to _____.
 (1) 0.0055 u (2) 1.00727 u
 (3) 1.6000 u (4) 1.672 u
2. Cathode rays are made up of _____.
 (1) positively charged particles
 (2) negatively charged particles
 (3) neutral particles
 (4) equal amount of positively and negatively charged particles
3. An α -particle is represented as _____.
 (1) ${}_1^4\alpha$ (2) ${}_2^4\alpha$
 (3) ${}_2^4\alpha$ (4) ${}_2^4\alpha$
4. The mass number and atomic number of an element are A and Z respectively. The number of neutrons in the atom of that element is calculated as _____.
 (1) A (2) Z
 (3) A + Z (4) A - Z
5. The nuclide having 6 protons and 8 neutrons is represented as _____.
 (1) ${}_6^{14}\text{X}$ (2) ${}_6^{14}\text{X}$
 (3) ${}_6^{14}\text{X}$ (4) ${}_6^{14}\text{X}$
6. The number of protons, neutrons and electrons in a neutral atom of the element ${}_{10}^{24}\text{X}$ are _____ respectively.
 (1) 89, 89, 242 (2) 89, 142, 89
 (3) 89, 71, 89 (4) 89, 231, 89
7. Which of the following pairs represent isobars?
 (1) ${}_2^4\text{He}$ and ${}_2^4\text{He}$ (2) ${}_{12}^{24}\text{Mg}$ and ${}_{12}^{24}\text{Mg}$
 (3) ${}_{19}^{40}\text{K}$ and ${}_{20}^{40}\text{Ca}$ (4) ${}_{19}^{40}\text{K}$ and ${}_{19}^{39}\text{K}$
8. Rutherford's model of an atom is unstable because _____.
 (1) the proton is an unstable particle
 (2) the electron is an unstable particle
 (3) the electron has very high kinetic energy
 (4) an accelerated electron in an orbit emits radiation
9. The SI unit of frequency of radiation is
 (1) cm^{-1} (2) Hertz
 (3) nm (4) metre
10. _____ is the number of wavelengths per unit length.
 (1) Amplitude (2) Velocity
 (3) Wavenumber (4) Frequency
11. The value of Rydberg constant for hydrogen atom is _____.
 (1) 109677 m^{-1} (2) 109677 cm^{-1}
 (3) $109677 \times 10^7 \text{ cm}^{-1}$ (4) $109677 \times 10^{-7} \text{ cm}^{-1}$
12. The emission spectrum of hydrogen in the ultraviolet region of electromagnetic radiation forms _____ series.
 (1) Pfund (2) Lyman
 (3) Balmer (4) Brackett
13. The spectral line obtained when an electron jumps from $n = 6$ to $n = 3$ level in hydrogen atom belongs to the _____ series.
 (1) Balmer (2) Lyman
 (3) Paschen (4) Pfund
14. Which of the following is INCORRECT regarding different series observed in hydrogen spectrum?
 (1) For Balmer series, $n_1 = 2$.
 (2) For Brackett series, $n_1 = 4$.
 (3) For Lyman series, $n_1 = 1$.
 (4) For Pfund series, $n_1 = 3$
15. According to Bohr's model for the hydrogen atom, when electrons revolve in the stationary orbits, _____.
 (1) the energies of the electrons remain constant
 (2) they become stationary
 (3) they gain potential energy
 (4) there is increase in energies of the electrons

Atom Structure Chemthink Answer Key

B Lingard



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