

**1** This question is about reversible reactions.

Sulfur trioxide is produced in a reversible reaction.

The equation for the reaction is:



**1.1** What is the correct balanced equation for the reaction?

Tick **one** box.

- ☐  $2\text{SO}_2 + \text{O}_2 \rightleftharpoons \text{SO}_3$
- ☐  $\text{SO}_2 + 2\text{O} \rightleftharpoons \text{SO}_3$
- ☐  $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$
- ☐  $2\text{SO}_2 + 2\text{O}_2 \rightleftharpoons 2\text{SO}_3$

[1 mark]

**1.2** The reaction of sulfur dioxide and oxygen to form sulfur trioxide is exothermic.

Which of the following statements is true?

Tick **one** box.

- ☐ The activation energy for the reaction of sulfur dioxide with oxygen is negative.
- ☐ Sulfur trioxide is colder than sulfur dioxide and oxygen.
- ☐ Heating the reaction will cause more sulfur trioxide to form.
- ☐ The reaction of sulfur trioxide to form sulfur dioxide and oxygen is endothermic.

[1 mark]

# Aqa Chem 5 June 2013 Question Paper

**Rob King**



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