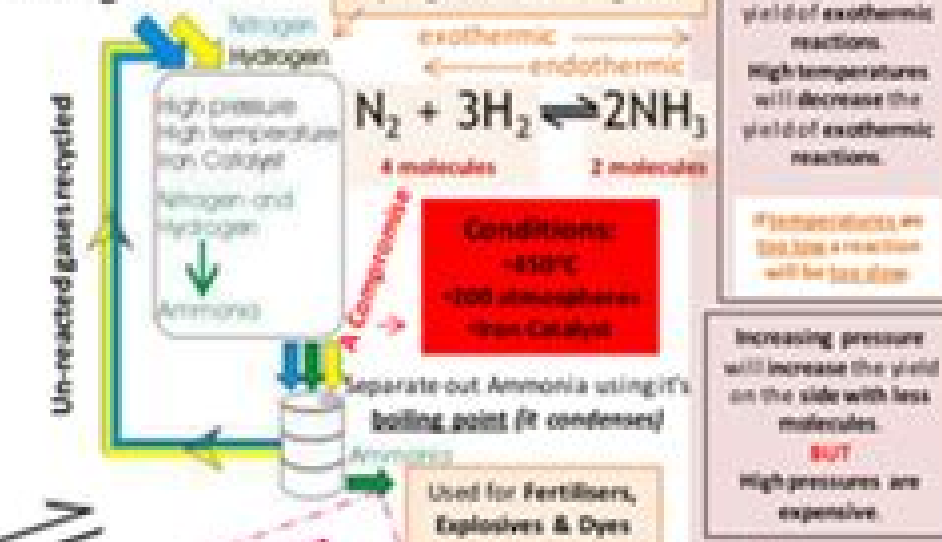


The Haber Process Making Ammonia



Reversible Reaction
One way will be exothermic, and the other will be endothermic.
In a closed system forward and backward reaction rates at equilibrium.

Flame Test

- Dip wire in sample
- Hold loop in flame
- Use the colour of the flame to identify the metal in the compound

Crystals (flame test):
Li⁺ (Red)
Ca²⁺ (Red)
Na⁺ (Yellow)
K⁺ (Lilac)
Ba²⁺ (Green)

Reactions with NaOH

- Add NaOH to sample
- Check to see what coloured precipitate (solid) forms.
- Check to see if adding more NaOH will make it dissolve

Unknown solution + NaOH

- White precipitate** → **dissolve** → **White** → **dissolve with more NaOH** → **Al³⁺**
- Light blue precipitate** → **dissolve with more NaOH** → **Mg²⁺ / Ca²⁺**
- Light green precipitate that starts to go brown** → **Fe²⁺**
- Reddish brown precipitate** → **Fe³⁺**

Titration

Use when you don't know the concentration of either acid or alkali

Acid of known concentration → you can use the stop-clock to add exactly the right amount

Pipette → used to measure 25cm³ of unknown alkali

Indicator → The reaction is finished when the colour changes

Acids and alkalis neutralise each other → You can use an indicator to show when this has happened. This is the **END POINT**.

Repeat three times until you can make an average

Calculation

Known Acid	Conc. of Unknown Alkali
Concentration (mol/dm ³)	0.5
Volume (dm ³)	50/1000 = 0.05
Moles (n)	0.5 × 0.05 = 0.025

Formula:
Concentration (mol/dm³) = Moles / Volume (dm³)

Example:
0.5 = 0.025 / Volume
Volume = 0.025 / 0.5 = 0.05 dm³

Acid Strength?

Two acids that are the same concentration can and be different strengths.
Strong acids dissociate (break) more in water completely.
Weak acids do not completely dissociate, not losing all their H⁺ ions.

Homologous series - a group of compounds which share a functional group giving them similar properties.

- All alcohols have the functional group -OH**
- All carboxylic acids have the functional group -COOH**
- All esters have the functional group -COO-**

Forensic gel electrophoresis

Chemical analysis can be used to ID individuals at crime scenes, pollution in rivers and matching transplant patients to donors in medicine.

Crime scene	Prison	Donor	Match
1	2	3	4
5	6	7	8
9	10	11	12
13	14	15	16
17	18	19	20
21	22	23	24
25	26	27	28
29	30	31	32
33	34	35	36
37	38	39	40
41	42	43	44
45	46	47	48
49	50	51	52
53	54	55	56
57	58	59	60
61	62	63	64
65	66	67	68
69	70	71	72
73	74	75	76
77	78	79	80
81	82	83	84
85	86	87	88
89	90	91	92
93	94	95	96
97	98	99	100

Negative Ions

- 1) Carbonate ions CO₃²⁻ → Add acid → Look for bubbles and test w/ Litmus (CO₂ is made)
- 2) Sulfate ions SO₄²⁻ → Add acid → Add Barium chloride → Look for white precipitate
- 3) Halide ions Cl⁻ Br⁻ I⁻ → Add Nitric acid → Add Silver Nitrate → Look for coloured precipitate

Colours: Chloride ions = white, Bromide ions = cream, Iodide ions = pale yellow

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