

Solutions, acids, and bases

1. VOCABULARY REVIEW OF SOLUTIONS

Solute - what is **dissolved** in the solvent to form a solution.

Solvent - the substance that **dissolves** the solute, ordinarily a liquid when the solute and the solvent come together to form a **homogeneous mixture**.

Miscible - two mix in all proportions (none to be mixed) - liquids will pour each over

Immiscible - don't mix in any proportions (none to be mixed)

Examples of miscible-immiscible - ethanol-water, oil-water, oil-water, oil-water

Solute - **amount** of **being dissolved** in liquid

Insoluble - a substance (tends to be a solid) that **will not dissolve** in a solvent (an example is sand and water)

Suspension - looks like a **solution** when **stirred**, but **settles** a bit of water (suspensions). Also it is a **hetero** - gaseous mixture where the small solid particles of a substance are spread throughout a liquid without settling.

Unsaturated - **solids** with **solubility chart** when the temperature would be enough to make a substance dissolve but at a greater amount of **solute** - **solubility less than** **solubility** when up (on the graph)

Saturated - **solids** with **solubility chart** when the temperature - **solubility relationship** is at the line in order words a **solution** with **solute** has **reached** at the **maximum possible amount**.

Supersaturated - contains **more than the max.** amount of **solute** that can be dissolved at the temperature on the solubility chart this would mean the grams of solute would go above the line.

2. SOLUBILITY VS. TEMPERATURE CHART



Above the line - **Supersaturated**

At the line - **Saturated**

Under the line - **Unsaturated**

3. METHODS OF SEPARATION

Filtration - separate particles by **size**

Distillation - separate by **boiling point** - **more** and **liquids** to be separated.

Decantation - allows one to remove a **liquid from a solid** by **transferring** the **liquid** to a **different** **beaker**.

Centrifugation - separating molecules by their **densities** by **spinning** them **quickly**.

Decanting - process of separating **insoluble** or **suspension** solutions by **pouring** away the **liquid** from the **solid**.

Chromatography - separation by the **rates** at which components of different solutions **move** through a **medium** (e.g. gel electrophoresis)

Acid And Bases Study Guide

**T. W. Graham Solomons, Craig B.
Fryhle, Scott A. Snyder**



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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