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Jay Storer



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Reports Announcements & Index ,1985-11

F&S Index International Annual ,1999

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation

$$8 3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8 3 \text{ mole } \text{H}_2\text{O}$$

 8 3 mole H_2S x 18g H_2O / 1 mole H_2O 149 4g H_2O is your theoretical yield so percent yield 137 1g H_2O actual yield 149 4g H_2O x 100
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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21 3 grams of O_2 produces a 47 3 yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21 3 \text{ grams } \text{O}_2}{32 \text{ g/mole}} = 0 66 \text{ moles } \text{O}_2$$

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web a write a balanced equation for the reaction

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

 4 al 3 o2 2 al2o3 b determine the theoretical yield of Al_2O_3 160 0

g al x 1 mol al x 2 mol al₂O₃ x 101 964 g al₂O₃ 302 3 g al₂O₃ 26 982 g al 4 mol al 1 mol al₂O₃ c determine the percent yield
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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 c is the limiting reagent 4 33 g of h 2 are left over 13

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 ch₃oh can be produced by the following reaction 2h₂ co ch₃oh calculate the theoretical yield of ch₃oh if 68 5 g of co is
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 what is the

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 \times 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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