

Bond Energy

What makes a reaction endothermic or exothermic?

Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** (ΔH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol _{rxn})		Reaction	Change in Energy (kJ/mol _{rxn})
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
 - When bonds are (broken/formed) there is a positive energy change.
 - Breaking bonds is (endothermic/exothermic).
 - When bonds are (broken/formed) there is a negative energy change.
 - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
 - How do the changes in energy for the reverse reactions compare?

The changes in energy for these reactions are the same, but opposite in sign.
 - Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.

Bond Energy Worksheet With Answers Pogil Chemistry

István Mayer



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Chemical Bonds and Bonds Energy R Sanderson, 2012-12-02 *Chemical Bonds and Bonds Energy* Second Edition provides information pertinent to the fundamental aspects of contributing bond energy and bond dissociation energy. This book explores the values that are useful in the interpretation of significant phenomena such as product distribution and reaction mechanisms. Organized into 12 chapters, this edition begins with an overview of the quantitative relationship among three basic properties of an atom: namely, nonpolar covalent radius, electronegativity, and homonuclear single covalent bond energy. This text then examines the quantitative means of evaluating the partial atomic charges that result from initial differences in the electronegativity of atoms that form a compound. Other chapters consider the recognition of the reduction of bond weakening not by multiplicity and in certain types of single covalent bonds. The final chapter deals with the application of the principal ideas and techniques to the oxidation of ethane. This book is a valuable resource for organic and inorganic chemists.

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[Comprehensive Handbook of Chemical Bond Energies](#) Yu-Ran Luo, 2007-03-09 Understanding the energy it takes to build or break chemical bonds is essential for scientists and engineers in a wide range of innovative fields including catalysis, nanomaterials, bioengineering, environmental chemistry, and space science. Reflecting the frequent additions and updates of bond dissociation energy (BDE) data throughout the literature, the *Comprehensive Handbook of Chemical Bond Energies* compiles the most recent experimental BDE data for more than 19 600 bonds of 102 elements. The author organizes the data by bond type, functional group, bond order, bond degree, molecular size, and structure for ease of use. Data can also be located using the Periodic table. The book presents data for organic molecules, biochemicals, and radicals as well as clusters, ions, hydrogen, and surface-bonded species, van der Waals complexes, isotopic species, and halogen clusters. It also introduces entirely new data for inorganics and organometallics. The final chapter summarizes the heats of formation for atoms, inorganic organic radicals, and monoatomic ions in the gas phase. The *Comprehensive Handbook of Chemical Bond Energies* offers quick access to experimental BDE data in the most inclusive, well-organized, and up-to-date

collection available today **Selected Solutions for Chemical Principles** Kenneth C. Brooks, Steven S. Zumdahl, 1992 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples Bond Orders and Energy Components, 2013 *Chemical Energetics and the Curriculum* Douglas James Millen, 1969 **CHEMICAL BONDS AND BOND ENERGY.** RT. SANDERSON, 1971

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Gray,1964 Bond Energies, Ionization Potentials, and Electron Affinities Vladimir Ivanovich Vedenev,1966 *Bond Orders and Energy Components* István Mayer,2016-10-14 While modern computational methods can provide us with the wave function of a molecule in numerical form most computer programs lack the sophisticated tools needed to extract chemical concepts from these wave functions Saving researchers vast time and potential confusion this volume collects and organizes those validated tools currently scattered throughout the literature and details their application It provides immediate access for those needing to calculate such critical factors as bond order and valence indices and atomic and diatomic contributions to molecular energy Supporting material is available for download from the authors continually updated website **Answer Guide for Chemistry, 2nd Edition** Leonard W. Fine,1978 **Chemical Bonding** M.S. Sethi & M. Satake,2010 Contents Chemical Bonding I Basic Concepts Chemical Bonding II Additional Aspects Intermolecular Force and Crystal Structures **Handbook of Bond Dissociation Energies in Organic Compounds** Yu-Ran Luo,2002-12-26 So many compounds so many experiments reported by so many researchers using so many methods Finding reliable data on bond dissociation energies BDEs can be like looking for a needle in a haystack But these data are crucial to work in chemical kinetics free radical chemistry organic thermochemistry and physical organic chemistry so where does one turn With data on almost 2900 bonds in 2500 organic compounds the Handbook of Bond Dissociation Energies in Organic Compounds is the first comprehensive collection of experimental BDE data Author Dr Yu Ran Luo has spent more than 10 years collecting assessing and tabulating BDE experiment results and presents them here in a well organized easy to use series of tables For each compound the tables list up to five recent experimental values indicate which is the recommended value give the method used to determine each value and provide references for each value An outstanding format and complete indexing both by compound name and by compound class make searching for data quick and easy While experimental BDE values carry inherent uncertainties the value of this compilation does not There is simply no other resource that provides as much data With this book when you need data on a particular bond dissociation energy you ll find it within just a few minutes Chemical Bonds Harry B. Gray,1994-12-05 This profusely illustrated book by a world renowned chemist and award winning chemistry teacher provides science students with an introduction to atomic and molecular structure and bonding This is a reprint of a book first published by Benjamin Cummings 1973 **Bonding Theory** Donald J. Royer,1968

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