

Why aren't the elements listed in alphabetical order?

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As charts go, the periodic table is a bit odd. It's not square. Large portions of the table appear to be missing at the top. It's not organized alphabetically so elements can be found easily. But to a chemist, the periodic table is a very powerful tool. The periodic table is organized by properties, both chemical and physical. Those properties relate to the electronic structure of the atoms of each element. In today's activity, you are going to study how the ground state electron configurations and structure of atoms are related to the shape and organization of the periodic table.

A blank periodic table grid with the following labels:

- s-block**: Located above the first two columns (Groups 1 and 2).
- p-block**: Located above the last six columns (Groups 13 to 18).
- d-block**: Located above the central block of ten columns (Groups 3 to 10).
- f-block**: Located above the two rows of seven columns at the bottom (Lanthanide and Actinide series).

The grid is organized into rows and columns. The first two columns are the s-block, the next ten columns are the d-block, the next six columns are the p-block, and the final two rows at the bottom are the f-block. The grid is currently empty, with only the labels and the grid lines visible.

1. Obtain a card with electron configurations for your group as assigned by your teacher.
 - a. Record your team number_____ and team name _____
 - b. For each element on your card, write the end of the configuration (the underlined portion).
2. What is similar about the last orbital notation appearing at the end of the configuration for each element in your set?

Breaking The Code Periodic Table Answers

Patrick Vollmar



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