

Unit 4: Bonding Review Sheet

1. Match the vocabulary term on the left with the definition on the right

A. Anhydrous Substance	1. _____ a molecular substance that dissolves in water to produce hydrogen ions (H^+)
B. Subscript	2. _____ Require Roman numerals in their names
C. Lewis Dot Structures	3. _____ tells how many electrons an atom gains or loses when forming a compound
D. Cations	4. _____ negatively charged ions formed by nonmetals
E. Polar Covalent	5. _____ a substance that is made up of molecules
F. Covalent Bond	6. _____ a group of atoms that are united by covalent bonds
G. Oxidation number	7. _____ represent valence electrons as dots around the element symbol
H. Acids	8. _____ hold atoms together in a metallic substance
I. Polyatomic ions	9. _____ positively charge ions formed by metals.
J. Chemical Bonds	10. _____ A compound that is composed entirely of ions
K. Empirical Formula	11. _____ one atom ions
L. Structural formula	12. _____ ionic compounds that absorb water into their solid structures
M. Molecule	13. _____ atoms gain, lose or share electrons to obtain a complete set of valence electrons
N. Nonpolar covalent	14. _____ compounds are formed when 2 atoms with different electronegativities form a covalent bonds
O. Hydrates	15. _____ water free substance
P. Transitions Metals	16. _____ to the lower right of an element - used to denote the number of atoms of that element in the compound
Q. Molecular formula	17. _____ compounds are formed when atoms with similar electronegativities share electrons
R. Anions	18. _____ contain ions of only 2 elements
S. Binary Ionic Compounds	19. _____ are the forces that link atoms together
T. Metallic Bonds	20. _____ specifies which atoms are bonded to each other in a molecule
U. Monatomic Ions	21. _____ denote the ratio of ions in a compound
V. Ionic Compounds	22. _____ ions formed by a group of elements
W. Molecular substance	23. _____ is formed when two elements that are not metals share a pair of electrons
X. Octet Rule	24. _____ Q tells how many atoms are in a single molecule of a compound

2. Describe the properties of ionic, covalent and metallic bonds. Write the formulas for the following polyatomic ions

- | | |
|--------------|----------------|
| A. Nitrate | E. Chlorate |
| B. Sulfate | F. Phosphate |
| C. Carbonate | G. Hydroxide |
| D. Ammonium | H. Bicarbonate |
| E. Chlorite | I. Acetate |
| F. Phosphite | J. Sulfite |
| | K. |

Bonding Vocabulary Review Sheet

Michal Rosen-Zvi

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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