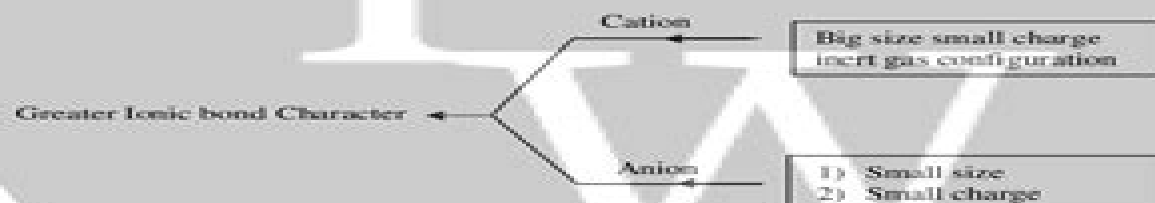


## CHAPTER

## 04

## CHEMICAL BONDING &amp; MOLECULAR STRUCTURE

- An atom (or) group of atoms having independent existence in nature is defined as molecule.
- Force of attraction that holds two atoms (or) oppositely charged ions in molecule is chemical bond.
- An atom must have eight electrons in its outermost energy level for its stability is called as octet rule.
- The number of electrons an atom loses (or) gains in an ionic bond formation is called electrovalency.
- According to Fajan's Rules
  - The no. of electrons an atom of the element loses (or) gains is known as electrovalency of that element.
  - Favourable conditions for cation formation.
    - 1) Low I.P
    - 2) Low charge on ion
    - 3) Big atomic size
    - 4) Inert gas configuration
  - Favourable conditions for anion formation.
    - 1) High E.N & E.A
    - 2) Small atomic size
    - 3) Low charge on ion



- $\text{LiCl}$  is more covalent in nature and so dissolved in non-polar solvents like alcohol (or) ether.
- Ions which have 18 electrons  $[ns^2np^6nd^{10}]$  in their outermost main energy level are stable and it is called as pseudo inert gas configuration (or) nickel group configuration.
- The energy liberated when an ionic crystal is formed from an assembly of isolated gaseous ions is called lattice energy.
- Lattice energy
 
$$(\cup) = \left[ -N_0 \frac{AZ^+Z^-e^2}{r} + N_0 \frac{Be^2}{r^9} \right]$$
  - $$\cup = \frac{AZ^+Z^-e^2N_0}{\epsilon_0} \left[ 1 - \frac{1}{n} \right]$$

(Born – Lande equation)

Coordination number of  $\text{Na}^+$  and  $\text{Cl}^-$  in  $\text{NaCl}$  crystal is (6:6)

Coordination number of  $\text{Cs}^+$  and  $\text{Cl}^-$  in  $\text{NaCl}$  crystal is (8:8)
  - The total amount of energy change in a reaction remains the same whether the reaction take place in one (or) several stages is known as Hess law of constant heat summation.

# Bonding Chemical Formulas

**Karin Nielsen-Saines**



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