

Changes in Bonding

Why haven't we progressed on regarding all of the elements we've met? Because most molecules are compounds. They have elements and stable enough to a net gain. Instead they must make other elements to form compounds. Why?

Shared Bonds—electrons shared to achieve a full set of valence electrons. This is usually 8 (what are the exceptions). Atoms can make stable molecules that meet this.

To fulfill the octet rule atoms will form gases, or share electrons. Atoms will gain electrons to achieve the full octet, usually with low electronegativity, non-metallic elements.

Why? Elements have a positive value than element with gas, metal, have low electronegativity. The elements don't have enough attraction for the electrons. (Electrons) have more positive and less electronegative elements in the energy level.

Ionization Energy

Ionization energy is the energy required to remove an electron from an atom. The first ionization energy is the energy required to remove the first electron. The second ionization energy is the energy required to remove the second electron.



Charge—positive ions (cations).
 Levels of electrons.
 Atoms—negative ions (anions).
 Levels of electrons.

How the ionization energy changes with 10 groups and 18 electrons.

How the ionization energy changes with 10 electrons.

Why has a well the first electron removed?
 Why?



Ion Bonds—Bonds of attraction

Ion bonds are formed between metal and non-metal atoms.



By the 10 groups
 a positive ion
 (cations) is formed.

To a
 form a new
 electron.

When, after a full set
 of 10 valence electrons,
 it has more a positive
 ion (cations).

Electron is positive
 positive 1 valence
 electron is 10 valence
 ion (cations) for stable.

To a
 form a new
 electron.

When, after a full set
 of 10 valence electrons,
 it has more a positive
 ion (cations) for
 charge (cations).



To form
 a stable
 compound.



Sodium Chloride—
 Table salt.

Why have we made compounds?

Why have we made compounds?

Mg^{2+} and Cl^{-} _____ Ca^{2+} and F^{-} _____

Ba^{2+} and O^{2-} _____ Sr^{2+} and S^{2-} _____

K^{+} and Br^{-} _____ Al^{3+} and O^{2-} _____

Be^{2+} and Cl^{-} _____ Mg^{2+} and O^{2-} _____

Bonding Answer Key For Stephen Murray

Sabine Zange



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