

Bond Energy

What makes a reaction endothermic or exothermic?

Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** (ΔH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol _{rxn})		Reaction	Change in Energy (kJ/mol _{rxn})
A)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
 - When bonds are broken (~~formed~~) there is a positive energy change.
 - Breaking bonds is endothermic (~~exothermic~~).
 - When bonds are broken (~~formed~~) there is a negative energy change.
 - Forming bonds is exothermic (~~endothermic~~).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
 - How do the changes in energy for the reverse reactions compare?
same absolute value, opposite sign
 - Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

amt of energy required to break is same as amount released when formed

Bond Energy Answer

CL Gary

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'CL Gary' text.

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General Chemistry Linus Pauling, 2014-11-24 Revised third edition of classic first year text by Nobel laureate Atomic and molecular structure quantum mechanics statistical mechanics thermodynamics correlated with descriptive chemistry Problems **Chemistry Homework for OCR A for Double and Separate Awards** Gareth Pritchard, 2001 Howework activites for OCR A Chemistry specifications **Essentials of Physical Chemistry 28th Edition** Bahl Arun/ Bahl B.S. & Tuli G.D., 2022-03 Essentials of Physical Chemistry is a classic textbook on the subject explaining fundamentals concepts with discussions illustrations and exercises With clear explanation systematic presentation and scientific accuracy the book not only helps the students clear misconceptions about the basic concepts but also enhances students ability to analyse and systematically solve problems This bestseller is primarily designed for B Sc students and would equally be useful for the aspirants of medical and engineering entrance examinations , **CUET PG Chemistry - Question Bank (2400 MCQs Chapterwise)** Mocktime Publication, 2400 MCQs CUET PG Question Bank Chemistry CUEG PG Previous Papers CUET PG Syllabus CUET PG Exam Pattern **Instrumental Methods of Chemical Analysis** Dr. B. K. Sharma, 1981 *CBSE (Central Board of Secondary Education) Class XI Science - Chemistry Topic-wise Notes | A Complete Preparation Study Notes with Solved MCQs* Mr. Rohit Manglik, 2023-07-07 EduGorilla s CBSE XI Science Chemistry Study Notes are the best selling notes for Class XI exams Their content is well researched and covers all topics related to CBSE XI Science Chemistry The notes are designed to help students prepare thoroughly for their exams with topic wise notes that are comprehensive and easy to understand The notes also include solved multiple choice questions MCQs for self evaluation allowing students to gauge their progress and identify areas that require further improvement These notes include Topics such as Some Basic Concepts of Chemistry Structure of Atom Chemical Thermodynamics Equilibrium Redox Reactions Organic Chemistry and Hydrocarbons These notes are perfect for understanding the pattern and type of questions asked by CBSE These study notes are tailored to the latest syllabus of CBSE XI Science Chemistry exams making them a valuable resource for exam preparation *Educart CBSE Class 11 Chemistry Question Bank 2026 (Strictly for 2025-26 Exam)* Educart, 2025-06-07 The Educart CBSE Class 11 Chemistry Question Bank 2026 is specially designed for students preparing for the 2025 26 session This book follows the latest CBSE syllabus and exam guidelines to help students build strong concepts and prepare well for their school exams Key Features 100% Based on Latest CBSE Syllabus Strictly follows the official CBSE Class 11 Chemistry syllabus for the 2025 26 academic year Chapterwise and Topicwise Questions Covers all chapters with a variety of CBSE type questions MCQs Very Short Short and Long Answer Assertion Reason and Case Based questions NCERT Focused Practice All questions are based on the NCERT Class 11 Chemistry textbook ensuring no confusion during school assessments Fully Solved Answers Includes complete step by step CBSE marking scheme solutions for all questions to help students learn how to write accurate answers in exams Competency Based Questions Questions framed to build understanding of real life

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Chapter-wise NCERT + Exemplar + Practice Questions with Solutions for CBSE Chemistry Class 11 - 2nd Edition Disha Experts, 2017-08-29 The book Chapter wise NCERT Exemplar Practice Questions with Solutions for CBSE Class 11 Chemistry has been divided into 3 parts Part A provides detailed solutions Question by Question of all the questions exercises provided in the NCERT Textbook Part B provides solutions to the questions in the NCERT Exemplar book Part C provides selected Practice Questions useful for the Class 11 examination along with detailed solutions The solutions have been designed in such a manner Step by Step that it would bring 100% Concept Clarity for the student

The Living Ocean Teacher's Guide , CBSE - ICSE Chemistry Part I Chandan Sengupta, First Publication October 2021 Place of Publication Arabinda Nagar Bankura 722101 This workbook will provide an ample scope in getting exposed to the system of acquiring skills and competence related to the understanding of chemistry It also exposes the student to the concepts of chemistry for enabling the aspirant in acquisition of skills related to chemistry Some of the worksheets are prepared along with model answers Some other worksheets are meant for self assessment and evaluation purposes It is also observed that some of the topics are specific to the referred curriculum Some other topics are varyingly incorporated in other streams of study Culmination of more than two streams will enable the fellow student to cope up with the preparatory works meant for Olympiads and other competitive examinations

Numerical Chemistry for Competitions Anu Sharma, 2009 An ideal book for the students of XI and XII CBSE ISC and the State Boards who are using Core Curriculum and also useful for the students preparing for various Engineering Medical Entrance Examinations

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Master the PCAT Peterson's, 2012-07-15 Peterson's Master the PCAT is an in depth review that offers thorough preparation for the computer based exam After learning about the structure format scoring and score reporting and the subtests and question types you can take a diagnostic test to learn about your strengths and weaknesses The next six parts of the eBook are focused on detailed subject reviews for each subtest verbal ability reading comprehension biology chemistry quantitative ability and writing Each review includes practice questions with detailed answer explanations You can take two practice tests to track your study progress The tests also offer detailed answer explanations to further improve your knowledge and understanding of the tested subjects The eBook concludes with an appendix that provides helpful information on a variety of careers in pharmacy and ten in depth career profiles

An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the

primary audience being undergraduate students and advanced high school students of chemistry

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Table of Contents Bond Energy Answer

1. Understanding the eBook Bond Energy Answer
 - The Rise of Digital Reading Bond Energy Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Bond Energy Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Bond Energy Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Bond Energy Answer
 - Personalized Recommendations
 - Bond Energy Answer User Reviews and Ratings
 - Bond Energy Answer and Bestseller Lists
5. Accessing Bond Energy Answer Free and Paid eBooks
 - Bond Energy Answer Public Domain eBooks
 - Bond Energy Answer eBook Subscription Services
 - Bond Energy Answer Budget-Friendly Options

6. Navigating Bond Energy Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Bond Energy Answer Compatibility with Devices
 - Bond Energy Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Bond Energy Answer
 - Highlighting and Note-Taking Bond Energy Answer
 - Interactive Elements Bond Energy Answer
8. Staying Engaged with Bond Energy Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Bond Energy Answer
9. Balancing eBooks and Physical Books Bond Energy Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Bond Energy Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Bond Energy Answer
 - Setting Reading Goals Bond Energy Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Bond Energy Answer
 - Fact-Checking eBook Content of Bond Energy Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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