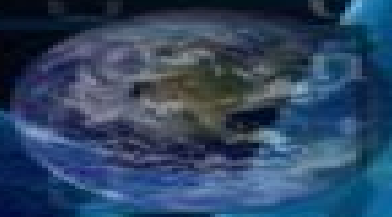


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 time the total pressure $p_a = x_a p_{\text{total}}$ mole fraction $x_a = \frac{\text{moles of substance a}}{\text{total moles}}$
 a is often denoted by x_a and

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 by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction
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 mass $C_6H_6 = 78 \text{ g/mol}$ hence C_6H_6 's number of moles $\frac{30}{78} = 0.3846 \text{ mol}$ molar mass of carbon tet $CCl_4 = 154 \text{ g/mol}$

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 g of water calculate the molarity and molality of the solution given $n(H) = 1$ $n(C) = 12$ $n(O) = 16$ given mass of solute urea 0.211 g mass
 of solvent water 100 g to find molarity of solution solution

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 of $CH_3OH = \frac{0.1722}{0.1722 + 0.073} = 0.703$ to solve more examples on mole fraction formula and practice more
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