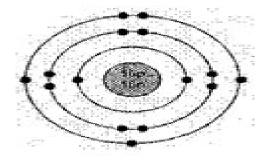
IV. Practice Questions

Section 4.1.

Atomic theory explains the formation of compounds: Atomic Theory and Bonding

Circle the letter of the best soower. You can use the Periodic Table of the Elements on page 2 of your Data Pages to help you.

- Which substomic particles are found in the nucleus of atoms?
 - A only protons
 - B. only electrons
 - C. protons and neutrons
 - D. protons and electrons
- When forming ions, elements on the left side of the periodic table tend to behave in which of the following ways?
 - A. Jose protons
 - D. gain protons
 - C. Tose electrons
 - D. guin electrons
- What is the name of the reactive family of elements that form 1 – loos?
 - A. halogens
 - R. noble gases
 - C. alkali metals
 - D. alkaline earth metals
- What is the atomic number of an atom with 19 protons, 19 electrons, and 21 neutrons?
 - A. 19
 - B. 21
 - C. 38
 - D. 39
- Which of the following describes an ion of chlorine?
 - A. 20 neutrons and 17 electrons
 - B. 20 neutrons and 18 electrons
 - C. 37 neutrons and 17 electrons
 - D. 37 neutrons and 18 electrons



- Identify the atom shown in the Bohr diagram above.
 - A. galliom
 - B. solfor
 - C. oneveren
 - D. phosphorus
- How many lone pairs and bonding pairs of electrons appear in the Lewis diagram shown below of water?

- A. two ione pairs and no bonding pairs
- B. two lone pairs and two bonding pairs
- C. no lone pairs and two bonding pairs
- D. four lone pairs and four bonding pairs

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BM King

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