



# Asus P5kc Manual

**Y Pai**



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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example  $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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 exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 va  
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 measure an appropriate volume of each liquid eg 25 cm 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less  
 exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of  
 sodium hydroxide solution with hcl as 57.9 kJ mol<sup>-1</sup>

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 produce salts neutralisation is the reaction between an acid and a base

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic  
 eg 1 hcl aq + naoh aq → nacl aq + h<sub>2</sub>o l for which ΔH = 57 kJ mol<sup>-1</sup> eg 2 h<sub>2</sub>no<sub>3</sub> aq + 2koh aq → 2h<sub>2</sub>o l for which ΔH = 57 kJ mol<sup>-1</sup>

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react

together to produce one mole of water for example  $\text{hcl aq naoh aq nacl aq h}_2\text{o aq}$  you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts

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