

Atoms & The Periodic Table

Unit 3 Test Study Guide

Lesson 1: The Atom

1. Who was the ancient Greek philosopher that came up with the idea that all things are made of small "uncuttable" particles?
2. What is an atom?
3. What was John Dalton's contribution to the atomic theory?
4. What are the parts of an atom?
 - a. Where can each of these particles be found within an atom?
 - b. How do the electrical charges of protons, neutrons, and electrons compare to each other?
 - c. How do the masses of protons, neutrons, and electrons compare to each other?
5. What is an atomic number?
 - a. What does this tell about the structure of an atom?
6. What is the atomic mass (mass number)?
 - a. What does this tell about the structure of an atom?
 - i. How can the atomic mass help you find the number of neutrons in an atom?

Lesson 2: The Periodic Table

1. What are elements?
2. What is the Periodic Table of Elements?
 - a. How are elements arranged on the Periodic Table?
3. Information found on each square of the Periodic Table:
 - a. Atomic Number (*See Lesson 1*)
 - b. What is a chemical name?
 - i. Examples
 - c. What is a chemical symbol?
 - i. Examples
 - d. Atomic Mass (*See Lesson 1*)
4. What are the 3 types of elements?
 - a. Where can these elements be found on the Periodic Table?
 - b. What are the characteristics of these types of elements?
5. What is a group?
 - a. What do the elements in a group have in common?
6. What is a period?
 - a. What do the elements in a period have in common?

Lesson 3: Electrons and Chemical Bonding

1. What is a chemical bond?
 - a. Why do atoms form bonds?
 - i. How can you figure out what an atom will do to become "happy"?
2. What is a group of atoms that are chemically bonded called?
3. What are valence electrons?

Atoms Periodic Table Study Guide Answer

Kerry L. Hull, Barbara Janson Cohen



Atoms Periodic Table Study Guide Answer:

Organic Chemistry, Student Study Guide and Solutions Manual David R. Klein, 2017-01-04 This is the Student Study Guide and Solutions Manual to accompany Organic Chemistry 3e Organic Chemistry 3rd Edition is not merely a compilation of principles but rather it is a disciplined method of thought and analysis Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems Readers must learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry Existing textbooks provide extensive coverage of the principles but there is far less emphasis on the skills needed to actually solve problems

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end of each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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