



Equations for AQA Physics

Triple Award

Recall and apply equations

You will have to learn each of the following equations off by heart for each exam paper, and know when and how to use them. Equation 9 is required for the Higher Tier papers only (as indicated by the symbol **HT**).



EQUATION	WORD EQUATION	SYMBOL EQUATION	
1	weight = mass $\times$ gravitational field strength (g)	$W = m g$	
2	work done = force $\times$ distance (along the line of action of the force)	$W = F d$	
3	force applied to spring = spring constant $\times$ extension	$F = k e$	
4	moment of a force = force $\times$ distance (normal to direction of force)	$M = F d$	
5	pressure = $\frac{\text{force normal to surface}}{\text{area of surface}}$	$p = \frac{F}{A}$	
6	Distance travelled = speed $\times$ time	$s = v t$	
7	acceleration = $\frac{\text{change in velocity}}{\text{time taken}}$	$a = \frac{\Delta v}{t}$	
8	Resultant force = mass $\times$ acceleration	$F = m a$	
9 HT	Momentum = mass $\times$ velocity	$p = m v$	
10	Kinetic energy = $0.5 \times \text{mass} \times (\text{speed})^2$	$E_k = \frac{1}{2} m v^2$	
11	Gravitational potential energy = mass $\times$ gravitational field strength (g) $\times$ height	$E_p = m g h$	
12	power = $\frac{\text{energy transferred}}{\text{time taken}}$	$P = \frac{E}{t}$	

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Peter Edmunds



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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations

requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction

therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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$\text{K}_2\text{PO}_4 + \text{Al}(\text{OH})_3 + \text{H}_2\text{O} \rightarrow \text{AlPO}_4 + \text{H}_2\text{PO}_4^- + \text{HPO}_4^{2-} + \text{H}_2\text{O}$
 $\text{K}_2\text{PO}_4 + \text{Al}(\text{OH})_3 + \text{H}_2\text{O} \rightarrow \text{AlPO}_4 + \text{H}_2\text{PO}_4^- + \text{HPO}_4^{2-} + \text{H}_2\text{O}$
 $\text{K}_2\text{PO}_4 + \text{Al}(\text{OH})_3 + \text{H}_2\text{O} \rightarrow \text{AlPO}_4 + \text{H}_2\text{PO}_4^- + \text{HPO}_4^{2-} + \text{H}_2\text{O}$

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds cl aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh₄ aq cr₂o₇²⁻ aq ag₂cr₂o₇ s 2nh₄ aq 2f aq because two nh₄ aq

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