

Cheboygan Hockey Association



Players completing the
Challenge will receive a
Trophy

5000 PUCK CHALLENGE

TRACKING SHEET

GOAL: Increase shot speed, quickness and accuracy

The Plan: Shoot 5000 pucks in 10 weeks

Weekly Goal: 500 shots per week.

Daily Goal: 100 shots per day (5 days per week)

WEEK 1:

		Date	Shots Total	Parent Initial
Day 1	100 Wrist Shot Low Corner			
Day 2	100 Wrist Shot Low Corner			
Day 3	100 Wrist Shot Low Corner			
Day 4	100 Wrist Shot Low Corner			
Day 5	100 Wrist Shot Low Corner			

WEEK 2:

		Date	Shots Total	Parent Initial
Day 1	100 Wrist Shot Top Corner			
Day 2	100 Wrist Shot Top Corner			
Day 3	100 Wrist Shot Top Corner			
Day 4	100 Wrist Shot Top Corner			
Day 5	100 Wrist Shot Top Corner			

WEEK 3:

		Date	Shots Total	Parent Initial
Day 1	100 Backhand Shot Low Corner			
Day 2	100 Backhand Shot Low Corner			
Day 3	100 Backhand Shot Low Corner			
Day 4	100 Backhand Shot Low Corner			
Day 5	100 Backhand Shot Low Corner			

5000 Puck Challenge Tracking Sheet

Lacey McSkate

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the Lacey McSkate text bar.

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web 1 answer a explanation the value of ΔH is negative therefore the reaction is exothermic or heat releasing 2 answer a 1
explanation the entire reaction was multiplied by 2 therefore ΔH must also be halved 3 answer c explanation option c is
incorrect because 566 kJ/mol is the ΔH per mole of O_2 and not for 2 moles

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web thermochemical equations equations that represent both mass and enthalpy changes $H_2O(s) \rightarrow H_2O(l)$ $\Delta H = 6.01 \text{ kJ/mol}$ this is
an endothermic process it requires 6.01 kJ to melt one mole of ice $H_2O(s)$ the enthalpy value will change if the number of
moles varies from the 1:1 reaction stoichiometry this is an exothermic process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy e of a system in any process is equal to the heat q added to the system plus the work w done on the system by its surroundings *e thermochemistry section review answer key prentice hall* - May 21 2022

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